

Organometallic Complexes. VII.* The Structure of the Iron Carbonyl Phenylacetylene Complex, $\text{Fe}_2(\text{CO})_6(\text{C}_6\text{H}_5\text{C}_2\text{H})_3$

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$\text{Fe}_2(\text{CO})_6(\text{C}_6\text{H}_5\text{C}_2\text{H})_3$ crystallizes with four molecules per unit cell in space group $P2_1/n$ with

$$a = 11.963 \pm 0.003, b = 20.442 \pm 0.005, c = 10.326 \pm 0.003 \text{ \AA}; \beta = 93^\circ 24' \pm 5'.$$

The structure has been determined from visually estimated three-dimensional intensity data obtained using $\text{Co K}\alpha$ radiation. The iron positions were found from a sharpened Patterson synthesis. Twenty-six of the light atoms were located by use of a minimum function and the remainder from two successive Fourier syntheses. Least-squares refinement led to a final R value of 15%.

The three phenylacetylene residues and one carbonyl group form a seven-membered carbon chain with phenyl groups at the 1, 3 and 6 positions and the carbonyl group in the 5 position. This chain is σ -bonded from the 1, 4 and 7 positions to one iron atom to form two fused five-membered rings with an iron and a carbon atom in common. This iron atom is also bonded to three carbonyl groups and is octahedrally coordinated. The second iron atom is linked to the remaining two carbonyl groups, to the three-electron donor system $\text{C}_1-\text{C}_2-\text{C}_3$ and to C_6-C_7 .

Bond lengths are: $\text{Fe}-\text{Fe}$, 2.501 ± 0.003 (e.s.d.); $\text{Fe}-\text{C}$ (carbonyl) 1.775 ± 0.012 ; $\text{Fe}-\text{C}(\sigma)$ 2.006, 2.097, 2.088; $\text{Fe}-\text{C}$ (three-centre) $1.96-2.23 \text{ \AA}$. All other bond lengths are normal.

Introduction

The reaction between iron dodecacarbonyl and phenylacetylene (Hübel & Braye, 1959) gives a number of organometallic compounds of which $\text{Fe}_2(\text{CO})_6(\text{C}_6\text{H}_5\text{C}_2\text{H})_3$ is of particular interest. Although very stable at ordinary temperatures, it decomposes at its melting point to give 1,3,5-triphenylbenzene. In boiling benzene it is converted to two isomers of triphenyltropone iron tricarbonyl, $\text{Fe}(\text{CO})_4(\text{C}_6\text{H}_5\text{C}_2\text{H})_3$. Recent work (Braye & Hübel, unpublished) has shown that it is also possible to obtain triphenyltropone from $\text{Fe}_2(\text{CO})_6(\text{C}_6\text{H}_5\text{C}_2\text{H})_3$ by treatment with triphenylphosphine. In view of these interesting reactions it seemed worth while to determine the structure of this complex.

Experimental

Crystals of the complex were obtained from a petroleum-ether/benzene solution as described by Hübel & Braye. They are thin black monoclinic plates lying on (010) and bounded by general forms.

Cell dimensions were determined by Farquhar & Lipson's (1946) method from oscillation photographs taken with $\text{Co K}\alpha$ radiation

$$(\lambda_{\alpha_1} = 1.78890, \lambda_{\alpha_2} = 1.79279 \text{ \AA})$$

and are

$$a = 11.963 \pm 0.003, b = 20.442 \pm 0.005, \\ c = 10.326 \pm 0.003 \text{ \AA}, \beta = 93^\circ 24' \pm 5'.$$

Systematic absences ($h0l$ for $h+l$ odd and $0k0$ for

k odd) were determined from Weissenberg photographs about the b and c axes. The space group is thus $P2_1/n$. This orientation was chosen to make the a and c axes as nearly orthogonal as possible. The observed density is $1.54 \pm 0.02 \text{ g.cm.}^{-3}$ and that calculated for $Z=4$ is 1.544 g.cm.^{-3} .

Multiple-film equi-inclination Weissenberg photographs were taken about the a ($h=0 \rightarrow 4$), b ($k=0$ and 1) and c ($l=0 \rightarrow 6$) axes using crystals $0.18 \times 0.13 \times 0.10 \text{ mm.}$, $0.11 \times 0.08 \times 0.08 \text{ mm.}$, and $0.12 \times 0.09 \times 0.07 \text{ mm.}$, respectively, the shortest dimension being in the b direction. In this way 3415 of the 3685 reflections accessible to $\text{Co K}\alpha$ radiation were recorded. Of these 2631 were strong enough to be observed. Intensities were estimated by visual comparison with a set of standard spots prepared from the same crystal. No attempt was made to correct for absorption or for variation in spot shape due to the irregular shapes of the crystals. Allowance was made for the extension and compression of spots on the upper-level Weissenberg photographs following Phillips (1954). Corrections for Lorentz and polarization factors were made by direct computation. Wilson's (1942) method was used to place the intensities on an absolute scale. This method also gave a mean temperature factor of 3.0 \AA^2 .

Determination of the structure

While three-dimensional intensity data were being collected, attempts were made to interpret the c - and a -axis projections. Sharpened Patterson functions were calculated for both these projections. In the

* Part VI. Krüerke, U., Hoogzand, C., & Hübel, W. (1961). *Chem. Ber.* **94**, 2817.

c-axis Patterson projection several more or less consistent sets of possible iron–iron vectors were found. All but one were eliminated by the use of a Buerger (1951) minimum function. A *c*-axis electron-density projection was calculated using signs from the iron-atom contributions alone but could not be interpreted in terms of any conceivable structure. It was impossible unambiguously to determine the *z* coordinates of the iron atoms from the *a*-axis Patterson projection, probably because these coordinates do not differ greatly and are both close to 0.75.

The three-dimensional intensity data were sharpened and then modified by the factor

$$(\sin \theta/\lambda)^4 \exp \{-25.64 (\sin \theta/\lambda)^2\}$$

(Shoemaker *et al.*, 1953) which has a maximum at $\sin \theta=0.5$ and a very small value at $\sin \theta=1$. From these modified data a three-dimensional Patterson function was calculated using the *M*-card method* (V. Schomaker, unpublished). The positions of the iron atoms were easily found from the Patterson function. The vectors between one iron atom and the three atoms related to it by the cell symmetry were used as the basic translations in the computation of a minimum function of rank 4. From the minimum function twenty-six light atoms were located in chemically reasonable positions. A retrospective examination showed clearly three more atoms in positions that were *a priori* unreasonable.

These twenty-eight atoms were used to calculate structure factors on an IBM 650 computer. The scattering factors used were taken from Berghuis *et al.* (1955) for C and O, from Freeman & Wood (1959) for Fe and, in the last stages of the work, from McWeeny (1951) for H. The values for Fe were reduced by a dispersion correction of 3.89 electron units calculated from the table given by James (1950). The discrepancy factor *R* was 0.41. A three-dimensional Fourier synthesis showed eight more atoms and a further structure-factor calculation (*R*=0.38) and Fourier synthesis indicated the positions of the two remaining atoms. A third calculation of structure factors (*R*=0.28) and Fourier synthesis confirmed these positions. The structure was refined by six applications of the method of least squares using the programme NYXR2 which is a development of NYXR1 (Friedlander, Love & Sayre, 1955) for the IBM 704 computer. The F_o were weighted simply according to the multiplicity with the exception that the non-observed data for which F_o was taken as $(\frac{1}{2}F_{\min}^2)^{\frac{1}{2}}$ were given 20% of this weight. Refinement of an isotropic temperature factor for each atom was introduced in the third cycle. The progress of the refinement is shown in Table 1 where

$$R = \Sigma(K|F_o| - |F_c|)/\Sigma K|F_o|$$

and $R' = \Sigma w(K|F_o| - |F_c|)^2$.

Table 1. *Progress of the least-squares refinement*

	<i>R</i>	<i>R'</i> 1st weighting	<i>R'</i> 2nd weighting
Parameters from 3rd Fourier synthesis	0.29	—	—
1st least-squares cycle	0.232	120516	—
2nd cycle	0.193	86053	—
3rd cycle, temperature factor refinement introduced	0.189	81815	—
4th cycle	0.178	72472	—
5th cycle	0.172	67961	—
6th cycle	0.169	66271	—
6th cycle, non-observed reflections omitted	0.151	64669	313650
7th cycle, hydrogen atoms included	0.145	—	299420
8th cycle	0.149	—	267106
9th cycle	0.152	—	261728

A three-dimensional difference synthesis was calculated using the structure factors calculated during the sixth refinement cycle for reflections of $\sin \theta/\lambda \leq 0.3 \text{ \AA}^{-1}$. There were more or less well-defined positive regions of height ca. 0.4 e. \AA^{-3} in the difference synthesis corresponding to the fifteen phenyl hydrogen atoms. In view of the broadness of some of these peaks these hydrogen atoms were placed on the diagonals of the benzene rings so that C–H=1.05 \AA .

The three remaining hydrogen atoms showed up as relatively sharp peaks from which their positions could be determined. There were about five other peaks of similar size which could not be accounted for by hydrogen atoms and in some cases might be due to anisotropic thermal vibration of phenyl carbon atoms. There were also negative regions (-0.7 e.\AA^{-3}) centred on the iron atoms; these are discussed below.

The refinement was completed by three more least-squares cycles in which hydrogen contributions were included and the non-observed reflections were omitted. The weighting system was changed to $w=500 p/F^2$ (where *p* is the multiplicity) with a minimum value of 1 and a maximum of 25. This change caused *R* to increase slightly (Table 1) owing to the change in scale factor. The effect of the hydrogen contributions to the structure factors is negligible and the improvement in *R'* during the last three refinement cycles was due mainly to the change in the weighting system. Examination of the values of observed and calculated structure factors from the ninth cycle showed that in most cases where $|F_o| - |F_c|$ was large this quantity was negative, suggesting a systematic error in scale or temperature factors. This was borne out by the negative regions round the iron atoms in the difference Fourier synthesis. It has recently been pointed out (Geller & Durand, 1960) that the programme NYXR2 has a theoretical fault in that it applies the scale factor correction to the observed data which should be kept constant in any least-squares treatment. They too found that this programme gave low values of the temperature factor.

* The master cards were kindly supplied by Dr V. Schomaker.

Table 2. Comparison of observed and calculated structure factors
N signifies a negative value

K	F _o	F _c	K	F _o	F _c	K	F _o	F _c	K	F _o	F _c	K	F _o	F _c	K	F _o	F _c	K	F _o	F _c
00	K 00	06 21 21	05 35 33 N	07 17 13 N	21 16 18 N	04 40 50	11 29 29	32 N	03 K 07											
02	101 112 N	08 17 28	06 40 39	08 44 48	22 10 12 N	05 29 33 N	12 46 55	52 N	00 39 40	00 32 36	01 32 36	02 23 23 N	03 26 28	04 27 28	05 27 28	06 26 28	07 27 28	08 27 28	09 27 28	10 27 28
04	43 N	09 33 38	08 50 52	09 24 27	02 K 01	04 46 51	11 22 25	51 N	05 11 14	06 12 14	07 12 14	08 12 14	09 12 14	10 12 14	11 12 14	12 12 14	13 12 14	14 12 14	15 12 14	
06	82 N	10 23 22 N	09 10 11	10 14 15	11 20 21	01 12 13	12 20 21	52 N	02 12 13	03 12 13	04 12 13	05 12 13	06 12 13	07 12 13	08 12 13	09 12 13	10 12 13	11 12 13	12 12 13	
08	25	11 39 42 N	10 14 15	11 14 15	12 20 21	01 12 13	12 20 21	53 N	02 12 13	03 12 13	04 12 13	05 12 13	06 12 13	07 12 13	08 12 13	09 12 13	10 12 13	11 12 13	12 12 13	
10	89	12 26 28 N	11 15 16	12 15 16	13 21 22	01 12 13	12 20 21	54 N	02 12 13	03 12 13	04 12 13	05 12 13	06 12 13	07 12 13	08 12 13	09 12 13	10 12 13	11 12 13	12 12 13	
12	38	14 17 13	13 15 16	14 17 13	15 20 21	01 12 13	12 20 21	55 N	02 12 13	03 12 13	04 12 13	05 12 13	06 12 13	07 12 13	08 12 13	09 12 13	10 12 13	11 12 13	12 12 13	
14	72 N	14 17 13	15 15 16	16 17 13	17 22 23	01 12 13	12 20 21	56 N	02 12 13	03 12 13	04 12 13	05 12 13	06 12 13	07 12 13	08 12 13	09 12 13	10 12 13	11 12 13	12 12 13	
16	8 N	15 17 2 N	15 15 16	16 17 13	17 22 23	01 12 13	12 20 21	57 N	02 12 13	03 12 13	04 12 13	05 12 13	06 12 13	07 12 13	08 12 13	09 12 13	10 12 13	11 12 13	12 12 13	
18	5	16 20 19	16 15 17	17 22 23	18 23 24	01 12 13	12 20 21	58 N	02 12 13	03 12 13	04 12 13	05 12 13	06 12 13	07 12 13	08 12 13	09 12 13	10 12 13	11 12 13	12 12 13	
20	8 N	18 20 19	18 17 18	19 22 23	20 23 24	01 12 13	12 20 21	59 N	02 12 13	03 12 13	04 12 13	05 12 13	06 12 13	07 12 13	08 12 13	09 12 13	10 12 13	11 12 13	12 12 13	
22	3 N	19 20 19	19 17 18	20 22 23	21 23 24	01 12 13	12 20 21	60 N	02 12 13	03 12 13	04 12 13	05 12 13	06 12 13	07 12 13	08 12 13	09 12 13	10 12 13	11 12 13	12 12 13	
00	K 01	00 K 06	01 16 17 N	02 16 17 N	03 14 15 N	04 14 15 N	05 14 15 N	06 14 15 N	07 14 15 N	08 14 15 N	09 14 15 N	00 22 16 N	01 22 16 N	02 22 16 N	03 22 16 N	04 22 16 N	05 22 16 N	06 22 16 N	07 22 16 N	
01	22 17	00 K 01	01 14 14 N	02 14 14 N	03 14 14 N	04 14 14 N	05 14 14 N	06 14 14 N	07 14 14 N	08 14 14 N	09 14 14 N	00 22 16 N	01 22 16 N	02 22 16 N	03 22 16 N	04 22 16 N	05 22 16 N	06 22 16 N	07 22 16 N	
03	80 4 N	01 42 39	00 56 58	01 46 53	02 51 58	03 52 53	04 52 53	05 52 53	06 52 53	07 52 53	08 52 53	00 22 16 N	01 22 16 N	02 22 16 N	03 22 16 N	04 22 16 N	05 22 16 N	06 22 16 N	07 22 16 N	
05	21 19	03 27 27	01 121 148 N	02 128 266	03 128 266	04 128 266	05 128 266	06 128 266	07 128 266	08 128 266	09 128 266	00 22 16 N	01 22 16 N	02 22 16 N	03 22 16 N	04 22 16 N	05 22 16 N	06 22 16 N	07 22 16 N	
06	78 82	05 30 30	03 55 56 N	04 55 56 N	05 55 56 N	06 55 56 N	07 55 56 N	08 55 56 N	09 55 56 N	00 22 16 N	01 22 16 N	02 22 16 N	03 22 16 N	04 22 16 N	05 22 16 N	06 22 16 N	07 22 16 N	08 22 16 N	09 22 16 N	
07	26 25 N	05 7 N	04 55 56 N	05 55 56 N	06 55 56 N	07 55 56 N	08 55 56 N	09 55 56 N	00 22 16 N	01 22 16 N	02 22 16 N	03 22 16 N	04 22 16 N	05 22 16 N	06 22 16 N	07 22 16 N	08 22 16 N	09 22 16 N	10 22 16 N	
09	55 4 N	07 52 52	06 55 56 N	07 55 56 N	08 55 56 N	09 55 56 N	00 22 16 N	01 22 16 N	02 22 16 N	03 22 16 N	04 22 16 N	05 22 16 N	06 22 16 N	07 22 16 N	08 22 16 N	09 22 16 N	10 22 16 N	11 22 16 N	12 22 16 N	
10	4 N	08 52 52	07 55 56 N	08 55 56 N	09 55 56 N	00 22 16 N	01 22 16 N	02 22 16 N	03 22 16 N	04 22 16 N	05 22 16 N	06 22 16 N	07 22 16 N	08 22 16 N	09 22 16 N	10 22 16 N	11 22 16 N	12 22 16 N	13 22 16 N	
11	4 2 N	09 52 52	08 55 56 N	09 55 56 N	10 55 56 N	11 55 56 N	12 55 56 N	13 55 56 N	14 55 56 N	15 55 56 N	16 55 56 N	17 55 56 N	18 55 56 N	19 55 56 N	20 55 56 N	21 55 56 N	22 55 56 N	23 55 56 N	24 55 56 N	
12	68 82 N	10 15 15	11 15 15 N	12 15 15 N	13 15 15 N	14 15 15 N	15 15 15 N	16 15 15 N	17 15 15 N	18 15 15 N	19 15 15 N	20 15 15 N	21 15 15 N	22 15 15 N	23 15 15 N	24 15 15 N	25 15 15 N	26 15 15 N	27 15 15 N	
14	15 11	11 12 26	12 13 15 N	13 14 15 N	14 15 15 N	15 16 15 N	16 17 15 N	17 18 15 N	18 19 15 N	19 20 15 N	20 21 15 N	21 22 15 N	22 23 15 N	23 24 15 N	24 25 15 N	25 26 15 N	26 27 15 N	27 28 15 N	28 29 15 N	
15	9 8	12 26 25	13 11 23	14 12 23	15 13 23	16 14 23	17 15 23	18 16 23	19 17 23	20 18 23	21 19 23	22 20 23	23 21 23	24 22 23	25 23 23	26 24 23	27 25 23	28 26 23	29 27 23	30 28 23
16	41 51	13 17 17	12 18 17	13 19 17	14 20 17	15 21 17	16 22 17	17 23 17	18 24 17	19 25 17	20 26 17	21 27 17	22 28 17	23 29 17	24 30 17	25 31 17	26 32 17	27 33 17	28 34 17	29 35 17
17	13 11 N	13 15 15	14 15 15	15 16 15	16 17 15	17 18 15	18 19 15	19 20 15	20 21 15	21 22 15	22 23 15	23 24 15	24 25 15	25 26 15	26 27 15	27 28 15	28 29 15	29 30 15	30 31 15	31 32 15
18	23 22 N	17 19 17	18 20 17	19 21 17	20 22 17	21 23 17	22 24 17	23 25 17	24 26 17	25 27 17	26 28 17	27 29 17	28 30 17	29 31 17	30 32 17	31 33 17	32 34 17	33 35 17	34 36 17	35 37 17
19	3 2 N	19 4	16 5 N	16 12 12	17 13 12	18 14 12	19 15 12	20 16 12	21 17 12	22 18 12	23 19 12	24 20 12	25 21 12	26 22 12	27 23 12	28 24 12	29 25 12	30 26 12	31 27 12	32 28 12
20	11 11 N	18 13 13	19 14 13	20 15 13	21 16 13	22 17 13	23 18 13	24 19 13	25 20 13	26 21 13	27 22 13	28 23 13	29 24 13	30 25 13	31 26 13	32 27 13	33 28 13	34 29 13	35 30 13	36 31 13
21	0 0	K 07	19 29 31 N	20 29 31 N	01 K 07	07 13 12 N	08 14 12 N	09 15 12 N	10 16 12 N	11 17 12 N	12 18 12 N	13 19 12 N	14 20 12 N	15 21 12 N	16 22 12 N	17 23 12 N	18 24 12 N	19 25 12 N	20 26 12 N	21 27 12 N
00	K 02	00 K 02	01 22 24 N	02 22 24 N	03 22 24 N	04 22 24 N	05 22 24 N	06 22 24 N	07 22 24 N	08 22 24 N	09 22 24 N	00 22 16 N	01 22 16 N	02 22 16 N	03 22 16 N	04 22 16 N	05 22 16 N	06 22 16 N	07 22 16 N	08 22 16 N
01	20 125 N	02 13 14 N	03 14 14 N	04 14 14 N	05 14 14 N	06 14 14 N	07 14 14 N	08 14 14 N	09 14 14 N	10 14 14 N	11 14 14 N	12 14 14 N	13 14 14 N	14 14 14 N	15 14 14 N	16 14 14 N	17 14 14 N	18 14 14 N	19 14 14 N	
02	70 71	05 24 24	06 24 24	07 24 24	08 24 24	09 24 24	01 K 02	02 K 02	03 K 02	04 K 02	05 K 02	06 K 02	07 K 02	08 K 02	09 K 02	10 K 02	11 K 02	12 K 02	13 K 02	14 K 02
03	62 70 N	05 21 21	06 21 21	07 21 21	08 21 21	09 21 21	01 K 02	02 K 02	03 K 02	04 K 02	05 K 02	06 K 02	07 K 02	08 K 02	09 K 02	10 K 02	11 K 02	12 K 02	13 K 02	14 K 02
04	35 35	07 21 21	08 21 21	09 21 21	10 21 21	11 21 21	01 K 02	02 K 02	03 K 02	04 K 02	05 K 02	06 K 02	07 K 02	08 K 02	09 K 02	10 K 02	11 K 02	12 K 02	13 K 02	14 K 02
05	45 42	07 21 21	08 21 21	09 21 21	10 21 21	11 21 21	01 K 02	02 K 02	03 K 02	04 K 02	05 K 02	06 K 02	07 K 02	08 K 02	09 K 02	10 K 02	11 K 02	12 K 02	13 K 02	14 K 02
06	54 56	08 21 21	09 21 21	10 21 21	11 21 21	12 21 21	01 K 02	02 K 02	03 K 02	04 K 02	05 K 02	06 K 02	07 K 02	08 K 02	09 K 02	10 K 02	11 K 02	12 K 02	13 K 02	14 K 02
07	54 56	09 21 21	10 21 21	11 21 21	12 21 21	13 21 21	01 K 02	02 K 02	03 K 02	04 K 02	05 K 02	06 K 02	07 K 02	08 K 02	09 K 02	10 K 02	11 K 02	12 K 02	13 K 02	14 K 02
08	34 35 N	10 21 20	11 21 20	12 21 20	13 21 20	14 21 20	01 K 03	02 K 03	03 K 03	04 K 03	05 K 03	06 K 03	07 K 03	08 K 03	09 K 03	10 K 03	11 K 03	12 K 03	13 K 03	
09	57 56 N	11 21 20	12 21 20	13 21 20	14 21 20	15 21 20	01 K 03	02 K 03	03 K 03	04 K 03	05 K 03	06 K 03	07 K 03	08 K 03	09 K 03	10 K 03	11 K 03	12 K 03	13 K 03	
10	19 13 N	00 K 09	11 21 19	12 21 19	13 21 19	14 21 19	01 K 04	02 K 04	03 K 04	04 K 04	05 K 04	06 K 04	07 K 04	08 K 04	09 K 04	10 K 04	11 K 04	12 K 04	13 K 04	
11	71 78	00 K 09	11 21 18	12 21 18	13 21 18	14 21 18	01 K 04	02 K 04	03 K 04	04 K 04	05 K 04	06 K 04	07 K 04	08 K 04	09 K 04	10 K 04	11 K 04	12 K 04	13 K 04	
12	38 39 N	01 24 24	02 24 24	03 24 24	04 24 24	05 24 24	06 24 24	07 24 24	08 24 24	09 24 24	10 24 24	11 24 24	12 24 24	13 24 24	14 24 24	15 24 24	16 24 24	17 24 24	18 24 24	
13	5 2 N	02 24 24	03 24 24	04 24 24	05 24 24	06 24 24														

Table 2 (cont.)

K	F ₀	F _c	K	F ₀	F _c	K	F ₀	F _c	K	F ₀	F _c	K	F ₀	F _c	K	F ₀	F _c			
13	54	53N	09	18	15N	07	14	10N	03	26	18	07	K 03		09	31	26N			
14	54	17	10	18	18	09	38	45	04	34	33	10	24	17	14	14	11			
15	20	10N	11	10	9	10	34	35N	05	37	36N	00	52	53	11	14	8			
16	24	17	12	11	11	12	12	10	06	35	28N	01	28	23N	12	18	11N			
17	19	15N	13	9	9	13	26	20N	07	24	17	03	24	18	13	36	31			
18	18	11N	14	13	21N	17	6	2N	08	35	29N	04	29	21N	14	17	7			
20	12	12N	18	10	15	10	10	08	28	26	05	39	34N	15	25	23N				
21	12	18	04	K 09	19	10	11	10	10	42	07	33	33N	16	9	5N				
04	K 03	03	7	7	05	K 05	15	21	13N	10	32	29	07	K 04	07	13	6N			
01	15	11	05	20	21N	00	35	35	18	14	12N	11	24	19	08	K 03	07			
02	22	20	09	10	12	00	29	26N	18	14	12N	13	26	19N	01	33	32			
03	21	25	10	10	9N	02	41	44N	14	46	42N	01	29	23N	11	23	16N			
04	15	10N	11	11	16N	03	31	29N	06	K 04	07	K 04	02	34	31N	03	32	28N		
05	62	62N	04	18	9	05	18	9	07	K 04	07	K 04	01	11	8N	04	35	26N		
06	18	15N	05	24	15	00	32	37N	01	15	10N	09	13	10	05	26	12N			
07	17	15N	04	K 10	06	38	37	01	23	17N	02	39	43N	10	10	3				
09	17	14N	00	17	21N	10	29	22N	03	20	16N	03	30	33	11	17	7			
10	12	01	12	1	1	11	43	41N	04	15	27	12	20	15N	02	35	29N			
11	18	14N	04	6	4	14	31	23	05	29	28N	04	29	29N	03	11	9			
12	40	40	06	8	9	13	19	9	06	22	14N	06	24	18	04	33	29			
13	19	18N	07	13	20	14	21	13N	07	58	58	16	10	3	06	11	7N			
14	31	30N	15	18	10	09	30	30	09	28	26	10	10	3	01	19	14N			
15	11	10	16	27	25	09	17	10	10	22	14N	10	12	9	02	20	13N			
17	31	28N	05	K 00	11	38	27N	11	27	28N	08	K 04	12	23	19N	03	14	9		
18	7	8N	02	26	8N	05	K 06	13	23	15N	15	14	9	04	26	11				
19	17	17	01	13	6N	05	K 06	14	18	15	16	20	22	05	14	15				
20	14	14	03	25	7N	01	40	37N	15	12	7	17	10	10N	06	24	24N			
04	K 04	05	51	52N	04	43	47	16	10	5N	07	K 05	10	10	7N	10	19	13N		
00	81	89	07	6	4N	08	22	16N	06	K 05	00	40	43N	13	26	24N	12	10	8	
01	45	52	08	6	6N	09	21	11N	02	12	8N	04	18	15N	11	11	6N			
02	21	18N	09	28	32	10	45	33	01	27	20N	01	12	7	09	K 05	11	K 02		
03	16	11N	10	26	25N	12	16	14N	02	34	30N	02	12	8N	05	16	8N			
04	51	52N	11	29	32N	15	22	21	03	30	27N	04	12	8N	06	18	15N			
05	50	50N	13	30	30N	04	28	22N	05	28	20N	09	K 06	02	16	8N				
07	51	54	13	11	12N	05	15	10	06	25	20	08	K 05	03	16	8N				
08	5	5N	14	27	20	05	K 07	06	19	12	07	26	22	07	32	29				
09	5	5N	15	24	22	05	K 07	07	25	20	08	26	22	07	32	29				
10	24	23	16	24	20	00	23	29	04	25	20	01	27	27N	04	12	10N			
11	21	24N	17	22	17N	01	44	52	09	52	44N	11	22	18	05	25	12			
12	18	14N	19	8	7	10	39	34N	12	25	18N	12	42	41	06	16	9N			
13	14	10	20	1	14N	11	27	19	13	18	15	09	34	30N	10	17	7N			
14	20	18N	05	K 09	12	25	17	14	14	6	06	19	10	07	26	16N				
15	26	23	13	19	18N	13	19	10	13	19	10	07	10	10	11	18	16N			
16	26	21	20	8	6N	15	34	31	06	K 08	07	K 09	09	28	23N	09	K 07	11	K 03	
17	9	8	17	18	12	00	28	31	01	15	23N	10	23	15N	00	21	28			
04	K 05	02	18	20N	06	K 00	06	K 06	01	22	15	13	22	16	01	14	17N			
01	32	82	02	18	20N	06	K 00	06	K 06	03	29	21N	08	K 05	02	16	8N			
02	25	24N	04	32	34	00	17	17	00	39	37N	05	20	13	03	19	11N			
03	83	98N	05	27	22N	01	89	105N	01	12	3N	06	25	17N	05	26	14N			
04	27	31N	01	15	12	02	16	13	03	17	10	09	37	37N	04	25	25N			
05	25	25	08	43	49N	03	27	28N	04	21	14	09	16	10N	01	42	39			
06	30	29	12	26	27N	04	19	16	05	17	9	10	48	30	17	05	7	4N		
07	13	8N	11	19	14N	06	5	12	06	29	24	11	28	17	05	20	20			
08	9	9	12	42	47	07	25	30N	08	10	20	04	20	13N	07	17	10			
09	21	21N	13	20	20	08	38	46	10	18	12N	07	K 07	05	26	24N				
10	12	8N	14	19	15	09	12	11	11	29	22	07	K 07	07	29	22N				
11	6	5	15	26	31N	10	17	15N	12	14	6N	00	12	10	09	26	24N			
12	31	27N	17	23	20N	11	23	25N	14	20	15	00	12	10	12	17	17N			
13	25	11	18	19	18N	13	15	10N	10	8	8	07	18	16N	11	19	13			
14	29	11	19	13	8N	14	8	4	06	48	56	07	K 09	05	25	18				
15	26	20	20	12	12	15	37	16	17	15	15	11	18	16N	13	14	8N			
16	26	20	17	20	14	12	42	4	4	4	15	12	7	5	14	3	10N			
17	17	10	18	13	11N	13	15	4	3	3	16	11	18	16N	14	3	8N			
18	22	24N	19	14	12N	14	32	31	07	K 01	07	K 01	09	K 01	04	21	20N			
17	5	7	20	12	12	15	25	16	19	13N	19	16	21	17	15	2	K 01			
04	K 07	05	K 03	06	K 02	01	44	52N	04	46	57	02	14	11	01	23	21N			
01	12	10	00	25	18N	00	26	29N	05	25	30	03	33	32	02	14	10N			
02	63	61	01	53	54N	02	61	73N	03	15	10N	01	56	68N	03	19	17N			
03	11	9	02	26	21	01	30	32	09	14	11N	02	14	10N	01	34	29			
04	8	10N	04	36	31	02	26	22N	10	38	39N	05	25	22	04	27	32			
06	8	5N	05	25	13N	03	33	36	14	31	25N	06	24	29	05	20	13			
08	7	8N	06	48	43N	04	37	36	14	31	25N	07	28	30N	01	29	23			
09	6	8N	07	30	24N	05	29	27N	18	17	15N	08	26	30N	06	29	24			
12	13	11	08	48	44	06	48	56	11	36	33N	09	17	15N	10	27	28N			
14	16	16	09	12	8	07	29	39	10N	12	19	13N	11	36	34	02	18	12N		
15	8	8N	10	27	21	08	34	35N	07	K 02	14	11	5N	14	18	14	04	22	20N	
16	16	26N	11	32	29	09	12	11	15	25	21	15	25	21	05	9	5	20N		
04	K 08	13	31	25N	10	19	11N	11	37	49	50	17	9	8N	09	33	33	07	17	21N
00	23	20	15	13	13N	05	K 04	16	8	5	07	20	12	00	23	21	06	16	13N	
01	13	13N	01	30	26	19	16	15	09	39	46N	01	34	37	02	21	16N	08	14	13N
02	10	7	02	26	24	12	16	15	12	34	25	04	27	32	03	28	31N	08	14	13N
03	10	3N	01	30	26	19	16	15	12	34	25	05	21	13N	05	21	15N	08	14	13N
05	8	8N	08	32	25	15	17	14	15	31	28	06	19	14N	06	21	15N	08	14	13N
06	10	11N																		

Table 2 (cont.)

K	F ₀	F _c	K	F ₀	F _c	K	F ₀	F _c													
12	K 0 4	0 6	1 2	7	1 4	1 3	1 2	0 2 N	K 0 7	0 3 N	K 0 3	0 3 N	K 0 9	1 0	1 1	1 0	0 5	1 5	1 4 N	0 4	1 5 32 N
0 0	4	3	0 7	7 2	7 9	1 5	9	8 N	0 1	4 2	4 8 N	0 0	5 1	5 0 N	1 1	2 2	2 1	0 4	1 2 52 N	1 32 N	
0 1	6	5	0 9	3 5	4 2 N	1 9	3 2	3 2	0 3	4 4	5	0 1	1 6	1 4 N	1 2	1 8	1 7	0 5	2 6 22	2 22	
0 2	9	6 N	1 0	2 8	2 4	2 0	9	9 N	0 3	3 5	4	0 2	7 6	6 N	1 3	2 2	2 2	0 7	2 2 21	21 N	
1 1	1	5 8	1 1	2 9	2 2 N	2 1	8	6	0 5	3 7	3 4	0 3	2 0	1 8	1 5	3 4	3 4	0 9	4 5 54 N	54 N	
1 2	2 9	2 2 N	2 2	3	4 N	0 5	3 2	3 4	0 5	4 9	4 7	0 5	2 5	2 8 N	1 5	3 7	3 1	1 0	3 4 39	39	
1 3	K 0 0	1 3	2 5	2 3	1 4	2 0	1 6 N	1 0	6	5	1 4	0 6	5 0	5 1 N	1 5	2 8	3 1	1 1	1 6 10 N	10 N	
0 1	1 2	1 4 N	1 5	2 3	1 7 N	0 2 N	K 0 2	1 0	5	6	0 8	1 3	5 N	0 8	1 4	1 1	1 2	1 8 35 N	35 N		
0 3	1 2	2 3	1 7	2 7	0 0	1 5 7	1 7 2 N	1 1	2 7	2 8 N	1 0	1 5	2 0	1 1	2 6	4	1 4	3 3	3 3	1 5 8 N	
1 3	K 0 1	1 8	5	6	0 1	1 5 7	1 0 4 N	1 3	2 2	2 4	1 8	1 2	7	1 6 N	1 2	3 3	2	1 6	2 4	2 1	
0 0	1 3	1 2	2 0	1 0	1 1	0 3	4 0	4 3 N	1 5	2 0	2 4	1 4	5 2	4 4 N	1 5	2 4	5	1 8	1 8	1 3	
0 1	1 6	2 4	0 1 N	K 0 6	0 6	0 5	8 9 N	1 6	4	3	1 5	1 8	1 3 N	1 5	5	6	2 0	8 N	0 5 N K 0 2		
0 1 N	K 0 1	0 1	1 0	1 1 N	0 8	3 7	3 2 N	1 7	1 0	8	1 7	1 0	8 N	0 3 N	K 0 10	0 5	1 7	1 2 N	0 1	2 2 1 N	
0 0	8 5	8 2 N	0 3	1 3	1 5	1 0	2 3	2 0 N	0 0	2 4	2 6 N	1 9	7	5 N	0 1	6	6 N	0 7	3 5 42 N	42 N	
0 1	1 8 6	1 9 0 N	0 4	6	2 N	1 1	4 5	4 9 N	0 0	2 4	2 6 N	2 1	1 2	1 1 N	1 0	2 4	3 0 N	0 5	2 1 18 N	18 N	
0 2	9 3	9 4	0 5	5 8	6 9	1 2	3 0	3 N	0 1	1 1	1 0 N	0 0	1 5	1 5 N	1 1	3 4	3 0 N	0 6	4 0 48 N	48 N	
0 3	1 0 6	1 0 9 N	0 5	1 9	1 9	1 3	1 8	1 7	0 2	2 4	2 4 N	0 6	1 6	1 6 N	1 2	1 2	1 2 N	0 7	3 1 33 N	33 N	
0 5	2 8 6	2 8 N	0 7	1 7	1 9 N	1 4	1 2	1 1 N	0 4	2 2	2 6 N	0 3 N	K 0 4	0 7	1 8	1 4 N	0 9	2 4 24 N	24 N		
0 6	2 8 2	2 8 N	0 8	1 3	1 4	1 5	1 3	8 N	0 4	2 2	2 6 N	1 7	1 0	8 N	1 5	2 1	2 1 N	1 0	16 11 N		
0 7	1 5 0	1 6 1 N	0 9	3 9	4 0 N	1 6	1 5	1 N	0 8	1 8	1 7 N	0 2	2 9	2 7 N	0 9	1 0	1 1	1 1 55 N	55 N		
0 8	3 6	3 6	1 1	2 4	2 5	2 0	1 8	2 N	1 2	2 1	1 8	0 3	4 2	3 9 N	0 3 N	K 11	1 5	2 4 9 11 N	11 N		
0 9	2 5	1 9	1 2	1 3	1 3 N	2 1	1 8	1 9 N	1 2	1 0	7	0 4	4 7	4 4 N	0 3 N	K 11	1 5	2 4 15 N	15 N		
1 0	3 4	3 7	1 5	6	6	2 2	6	1 0	1 7	2 0	0 5	3 7	3 3 N	0 4 N	K 11	1 5	2 4 15 N	15 N			
1 1	2 3	2 6 N	1 6	1 9	1 4	1 4	7	2 2 N	1 4	7	7	0 6	5	5 N	0 4 N	K 0 6	1 7	2 1 12 N			
1 2	2 3	2 1 N	1 7	1 2	1 0	0 2 N	K 0 3	0 1	2 3	2 0 N	0 0	4	4 N	0 0	3 0	2 9 N	1 8	1 9 13 N	13 N		
1 3	1 3	1 N	1 8	1 6	2 0 N	0 2	2 2	2 0 N	0 1	2 3	2 5 N	0 3	1 0	1 2 N	0 0	3 0	2 9 N	2 0	1 2 12 N		
1 4	1 6 4	1 6 4	0 1 N	K 0 7	0 3	3 6	3 6	0 2 N	K 0 9	1 0	1 8	1 8 N	1 1	4 7	5 3 N	0 3 N	K 0 13	0 4 20 13 N	13 N		
1 5	3 2	3 8	0 2	2 2	2 2 N	0 1	2 3	2 0 N	1 1	4 4	4 2 N	0 4 N	K 0 1	0 5	2 0	18	0 5 18 N	18 N			
1 6	2 0	2 3 N	0 3	1 8	1 4 N	0 4	1 0	1 0 5	0 3	2 5	2 5 N	0 5	2 0	2 0	0 5	2 0	18	0 5 18 N	18 N		
1 7	1 8	1 6	0 1 N	K 0 7	0 3	3 6	3 6	0 2 N	K 0 9	1 0	1 8	1 8 N	1 1	4 7	5 3 N	0 3 N	K 0 13	0 4 20 13 N	13 N		
1 8	7	4	0 2	1 5	1 8 N	0 4	1 0	1 0 5	0 3	2 5	2 5 N	0 4 N	K 0 1	0 5	2 0	18	0 5 18 N	18 N			
2 0	2 2	2 2	0 1	1 5	1 8 N	0 5	9 6	1 0 0	0 9	1 5	1 2 N	1 5	7	7	0 5	2 0	18	0 5 18 N	18 N		
0 1 N	K 0 2	0 5	1	1 1	1 1 N	0 8	4 1	1 2 N	1 2	1 4	1 6 N	1 3	1 0	1 2 N	0 5 N	K 0 5	0 5	2 0	18	0 5 18 N	
0 1	4 7	4 6 N	0 7	2 9	3 8 N	1 0	1 4	3 9 N	1 0	1 4	3 9 N	0 5 N	K 0 5	0 5	2 2	2 2 N	0 5	1 8 5 3 N	5 3 N		
0 2	4 2	4 2	0 9	5	7 N	1 1	1 8	1 3 N	0 2 N	K 0 6	1 2	4 4	4 2 N	0 4 N	K 0 1	0 5	2 0	2 0	0 5 18 N	18 N	
0 3	3 4	3 3 N	1 0	1 8	1 4 N	1 2	1 1	1 4 N	0 1	2 5	2 5 N	1 3	1 6	1 2 N	0 1	5 9 2 5 N	2 5 N	0 1 5 3 30 N	30 N		
0 4	9 5	1 0 7 N	1 1	5	7	1 3	1 6	1 1 N	0 0	2 1	2 1 N	0 3	2 2	2 2 N	0 1	5 9 2 5 N	2 5 N	0 1 5 3 30 N	30 N		
0 5	2 6	2 3 N	1 4	2 3	2 4	1 4	1 7	1 7 N	0 1	2 5	2 5 N	1 5	1 6	1 2 N	0 1	5 9 2 5 N	2 5 N	0 1 5 3 30 N	30 N		
0 6	4 5	4 6	1 5	1 0	7	1 5	1 0	7	1 6	2 0	1 4	1 4 N	2 0	1 1	1 2 N	0 1	5 9 2 5 N	2 5 N	0 1 5 3 30 N	30 N	
0 7	2 8	2 9	1 7	1 2	1 0 N	1 6	1 7	1 0 N	1 6	1 7	1 0 N	0 1	2 4	2 4 N	0 4 N	K 0 7	0 5	2 0 18 N	18 N		
0 8	9 8	1 0 2 N	1 7	2 5	2 0 N	0 4	1 6	3 4 N	0 5	2 3	2 4 N	0 5	2 3	2 4 N	0 4 N	K 0 7	0 5	2 0 18 N	18 N		
0 9	6 8	7 3 N	0 1 N	K 0 8	1 8	2 3	2 3 N	0 4	1 6	3 4 N	0 5	2 3	2 4 N	0 4 N	K 0 7	0 5	2 0 18 N	18 N			
1 0	6 4	6 4	0 1	1 0	3 N	1 9	2 3	2 3 N	0 5	2 3	2 4 N	0 5	2 3	2 4 N	0 4 N	K 0 7	0 5	2 0 18 N	18 N		
1 1	4 1	4 4	0 2	1 2	1 3 N	1 2	1 1	1 4 N	0 2 N	K 0 4	0 6	1 4	1 4 N	0 5 N	K 0 5	0 5	2 0 18 N	18 N			
1 2	2 1	1 8 N	0 2	3 2	3 9 N	1 2	1 1	1 4 N	0 2 N	K 0 4	0 6	1 4	1 4 N	0 5 N	K 0 5	0 5	2 0 18 N	18 N			
1 3	4 1	4 4	0 3	5 6	6 N	1 2	1 1	1 4 N	0 2 N	K 0 4	0 6	1 4	1 4 N	0 5 N	K 0 5	0 5	2 0 18 N	18 N			
1 4	1 8	1 6	0 4	1 7	1 9 N	1 2	1 1	1 4 N	0 2 N	K 0 4	0 6	1 4	1 4 N	0 5 N	K 0 5	0 5	2 0 18 N	18 N			
1 5	4	5	0 5	1 0	1 1 N	0 0	1 2	1 6 N	0 2 N	K 0 4	0 6	1 4	1 4 N	0 5 N	K 0 5	0 5	2 0 18 N	18 N			
1 6	7	4 N	0 6	2 1	2 4 N	0 1	1 2	1 6 N	0 2 N	K 0 4	0 6	1 4	1 4 N	0 5 N	K 0 5	0 5	2 0 18 N	18 N			
1 7	4	4 N	0 6	2 1	2 4 N	0 1	1 2	1 6 N	0 2 N	K 0 4	0 6	1 4	1 4 N	0 5 N	K 0 5	0 5	2 0 18 N	18 N			
1 8	1 6	1 6 N	0 7	1 7	1 8 N	0 1	1 2	1 6 N	0 2 N	K 0 4	0 6	1 4	1 4 N	0 5 N	K 0 5	0 5	2 0 18 N	18 N			
1 9	8	8 N	0 8	2 1	2 3 N	0 1	1 2	1 6 N	0 2 N	K 0 4	0 6	1 4	1 4 N	0 5 N	K 0 5	0 5	2 0 18 N	18 N			
2 0	2 0	1 9 N	0 8	2 1	2 3 N	0 1	1 2	1 6 N	0 2 N	K 0 4	0 6	1 4	1 4 N	0 5 N	K 0 5	0 5	2 0 18 N	18 N			
2 1	1 2	1 1 N	0 9	2 1	2 3 N	0 1	1 2	1 6 N	0 2 N	K 0 4	0 6	1 4	1 4 N	0 5 N	K 0 5	0 5	2 0 18 N	18 N			
0 1 N	K 0 4	0 5	1 0	1 0	1 0 N	1 3	1 2	1 6 N	0 2 N	K 0 6	0 5	1 0	1 0 N	0 1	5 9 48 N	48 N	0 1 5 3 30 N	30 N			
0 2	4 6	5 2	0 6	1 8	1 9 N	1 3	1 2	1 6 N	0 2 N	K 0 6	0 5	1 0	1 0 N	0 1	5 9 48 N	48 N	0 1 5 3 30 N	30 N			
0 3	4 9	5 2	0 6	1 8	1 9 N	1 3	1 2	1 6 N	0 2 N	K 0 6	0 5	1 0	1 0 N	0 1	5 9 48 N	48 N	0 1 5 3 30 N	30 N			
0 4	2 3	2 5 N	0 1 N	K 11	1 H	1 6	1 3	0 3 N	0 3 N	K 0 2	0 8	1 0	1 4 N	0 1	5 9 48 N	48 N	0 1 5 3 30 N	30 N			
0 5	1 3 9	1 5 6 N	0 0	2 6	3 0 N	2 0	3	3 N	0 1	1 0	1 2 N	1 1	9	1 0	4 0	3 5 N	0 4 N	K 10	0 5 27 31 N	31 N	
0 7	1 7	1 8	0 0	2 6	3 0 N	2 0	3	3 N	0 2	5 2	5 2 N	1 2	1 1	1 0	4 0	3 5 N	0 4 N	K 10	0 5 27 31 N	31 N	
0 8	4 6	5 5	0 2	8	9 N	1 3	1 2	1 6 N	0 2	5 2	5 2 N	1 3	1 1	1 0	4 0	3 5 N	0 4 N	K 10	0 5 27 31 N	31 N	
0 9	4 6	5 5	0 2	8	9 N	1 3	1 2	1 6 N	0 2	5 2	5 2 N	1 3	1 1	1 0	4 0	3 5 N	0 4 N	K 10	0 5 27 31 N	31 N	
1 0	8 2	8 8	0 3	1 5	4 N	0 2 N	K 0 6	0 5	1 0	1 2 N	1 2	1 1	1 0	4 0	3 5 N	0 4 N	K 10	0 5 27 31 N	31 N		
1 1	1 0	9	0 4	1 5	2 1	0 3 N	0 5	1 0	1 2 N	1 2	1 1	1 0	4 0	3 5 N	0 4 N	K 10	0 5 2				

Table 2 (cont.)

Using the structure factors from the ninth least-squares cycle, a two-by-two matrix was calculated on the IBM 602A machine for the refinement of scale and overall temperature factors with the F_o fixed. This resulted in a change of 3.5% in the scale factor in the opposite sense (reducing the F_c) to that calculated in the ninth 704 cycle, which reduced the F_o by 1.5% or, with the 'fudge' factor of 0.5 built into the programme, by 0.75%. The increase in the overall temperature factor was 0.09 Å². The effect of these changes on R and $R'/\sum w F_o^2$ was very small. Final values of observed and calculated structure factors are given in Table 2 and final atomic parameters together with standard deviations calculated from the least-squares matrix in Table 3.

An examination of the distribution (which is non-Gaussian) of the eighteen C-C bonds in the phenyl groups suggests a standard deviation of their bond lengths of 0.029 Å compared with a mean value of 0.019 Å calculated from the standard deviations of

atomic positions. The most pessimistic view based on the absolute equality of these eighteen bonds, when the variance would be due entirely to random errors, leads to the hypothesis that all standard deviation values in Tables 3, 4, and 5 should be multiplied by 1.5. While there is no scientific justification for this, as the method of calculating standard deviations by the inversion of the least-squares matrix is the most pessimistic known, it may be pointed out that such an increase would in no way affect the discussion that follows.

Discussion of the structure

The molecule, which is asymmetric, exists in two enantiomorphic forms and the substance crystallizes as the racemate. Bond lengths and bond angles are listed in Tables 4 and 5. Fig. 1 shows the molecule which has two distinct parts. The first is centred on Fe_{11} , which is octahedrally coordinated to three carbonyl groups and to the three atoms C_1 , C_4 and C_7 of the

Table 3. Final atomic parameters with their standard deviations

	x/a	y/b	z/b	B (Å ²)	$\sigma(x)$ (Å ²)	$\sigma(y)$ (Å ²)	$\sigma(z)$ (Å ²)	$\sigma(B)$ (Å ²)
Fe ₁	0.5410	0.3893	0.7315	2.96	0.002	0.002	0.002	0.04
Fe ₂	0.3858	0.3243	0.8202	2.81	0.002	0.002	0.002	0.04
O ₁	0.6923	0.2314	0.8467	4.17	0.009	0.008	0.008	0.16
O ₂	0.2594	0.4159	0.9705	6.52	0.013	0.011	0.012	0.24
O ₃	0.2590	0.2180	0.9307	6.17	0.012	0.011	0.011	0.23
O ₄	0.4994	0.5236	0.8205	6.44	0.012	0.011	0.011	0.24
O ₅	0.5693	0.4118	0.4536	6.29	0.012	0.011	0.011	0.24
O ₆	0.7786	0.3967	0.8070	6.28	0.012	0.011	0.011	0.24
C ₁	0.3781	0.3728	0.6545	3.28	0.013	0.010	0.011	0.19
C ₂	0.3647	0.3066	0.6192	3.23	0.012	0.010	0.010	0.19
C ₃	0.4507	0.2630	0.6662	3.08	0.012	0.010	0.010	0.18
C ₄	0.5663	0.2897	0.6942	2.79	0.011	0.010	0.010	0.17
C ₅	0.6069	0.2647	0.8242	2.71	0.011	0.009	0.010	0.17
C ₆	0.5408	0.2897	0.9292	3.14	0.012	0.010	0.010	0.19
C ₇	0.5221	0.3578	0.9130	3.52	0.013	0.011	0.011	0.20
C ₈	0.3083	0.3811	0.9056	4.11	0.014	0.012	0.013	0.23
C ₉	0.3083	0.2608	0.8860	4.38	0.015	0.012	0.013	0.24
C ₁₀	0.5150	0.4722	0.7820	4.08	0.014	0.012	0.012	0.23
C ₁₁	0.5586	0.4046	0.5627	4.10	0.014	0.012	0.013	0.23
C ₁₂	0.6848	0.3958	0.7784	4.58	0.015	0.013	0.013	0.26
C ₁₃	0.3086	0.4233	0.5852	3.61	0.013	0.011	0.011	0.21
C ₁₄	0.2572	0.4754	0.6480	4.91	0.016	0.014	0.014	0.27
C ₁₅	0.1872	0.5171	0.5775	5.74	0.018	0.015	0.016	0.31
C ₁₆	0.1703	0.5088	0.4477	5.77	0.018	0.016	0.016	0.31
C ₁₇	0.2154	0.4582	0.3818	6.08	0.018	0.016	0.016	0.33
C ₁₈	0.2844	0.4148	0.4513	4.14	0.014	0.012	0.012	0.23
C ₁₉	0.4305	0.1909	0.6515	2.97	0.012	0.010	0.010	0.18
C ₂₀	0.3221	0.1653	0.6391	4.22	0.015	0.012	0.012	0.24
C ₂₁	0.3105	0.0972	0.6120	4.46	0.015	0.013	0.013	0.25
C ₂₂	0.3997	0.0576	0.6053	4.88	0.016	0.014	0.014	0.26
C ₂₃	0.5074	0.0830	0.6202	4.54	0.015	0.013	0.013	0.25
C ₂₄	0.5248	0.1477	0.6445	3.85	0.013	0.012	0.012	0.21
C ₂₅	0.5323	0.2538	0.0519	3.09	0.012	0.010	0.011	0.19
C ₂₆	0.5050	0.2867	0.1668	4.26	0.014	0.012	0.013	0.24
C ₂₇	0.4970	0.2530	0.2793	5.54	0.018	0.014	0.016	0.30
C ₂₈	0.5144	0.1876	0.2892	5.04	0.016	0.014	0.014	0.28
C ₂₉	0.5404	0.1531	0.1768	4.69	0.015	0.013	0.014	0.26
C ₃₀	0.5518	0.1885	0.0600	4.37	0.015	0.012	0.013	0.24

Table 4. Interatomic distances

Bond	Length	Standard deviation	Bond	Length	Standard deviation
Fe ₁ -Fe ₂	2.501 Å	0.003 Å	C ₁₄ -C ₁₅	1.373	0.022
Fe ₁ -C ₁₀	1.806	0.012	C ₁₅ -C ₁₆	1.354	0.023
Fe ₁ -C ₁₁	1.796	0.013	C ₁₆ -C ₁₇	1.367	0.023
Fe ₁ -C ₁₂	1.764	0.016	C ₁₇ -C ₁₈	1.382	0.021
Fe ₂ -C ₈	1.756	0.013	C ₁₈ -C ₁₃	1.406	0.017
Fe ₂ -C ₉	1.754	0.013	C ₃ -C ₁₉	1.499	0.014
Fe ₂ -C ₁	2.088	0.013	C ₁₉ -C ₂₀	1.398	0.019
Fe ₁ -C ₄	2.097	0.010	C ₂₀ -C ₂₁	1.424	0.018
Fe ₁ -C ₇	2.006	0.011	C ₂₁ -C ₂₂	1.344	0.021
Fe ₂ -C ₁	1.975	0.011	C ₂₂ -C ₂₃	1.389	0.022
Fe ₂ -C ₂	2.108	0.011	C ₂₃ -C ₂₄	1.359	0.018
Fe ₂ -C ₃	2.203	0.011	C ₂₄ -C ₁₉	1.438	0.017
Fe ₂ -C ₆	2.227	0.012	C ₆ -C ₂₅	1.473	0.015
Fe ₂ -C ₇	1.965	0.013	C ₂₅ -C ₂₆	1.419	0.016
C ₈ -O ₂	1.159	0.017	C ₂₆ -C ₂₇	1.359	0.020
C ₉ -O ₃	1.167	0.017	C ₂₇ -C ₂₈	1.356	0.020
C ₁₀ -O ₄	1.142	0.016	C ₂₈ -C ₂₉	1.408	0.019
C ₁₁ -O ₅	1.151	0.017	C ₂₉ -C ₃₀	1.420	0.019
C ₁₂ -O ₆	1.144	0.020	C ₃₀ -C ₂₅	1.357	0.016
C ₁ -C ₂	1.407	0.015	Non-bonded approaches		
C ₂ -C ₃	1.425	0.016	Fe ₂ -C ₄	2.682	0.011
C ₃ -C ₄	1.499	0.016	Fe ₂ -C ₅	2.910	0.011
C ₄ -C ₅	1.490	0.014	C ₁ -C ₇	3.105	0.017
C ₅ -O ₁	1.238	0.014	C ₂ -C ₆	3.743	0.016
C ₅ -C ₆	1.471	0.014	C ₃ -C ₅	2.407	0.016
C ₆ -C ₇	1.419	0.015			
C ₁ -C ₁₃	1.433	0.017			
C ₁₃ -C ₁₄	1.408	0.018			

Table 5. Bond angles

	Angle	Standard deviation		Angle	Standard deviation
$\text{C}_1-\text{Fe}_1-\text{C}_4$	85.1°	0.4	$\text{C}_3-\text{C}_4-\text{C}_5$	107.2°	0.9
$\text{C}_1-\text{Fe}_1-\text{C}_7$	98.6	0.5	$\text{C}_4-\text{C}_5-\text{O}_1$	125.4	1.0
$\text{C}_4-\text{Fe}_1-\text{C}_7$	83.4	0.4	$\text{O}_1-\text{C}_5-\text{C}_6$	121.7	1.0
$\text{C}_1-\text{Fe}_1-\text{C}_{10}$	95.1	0.5	$\text{C}_4-\text{C}_5-\text{C}_6$	112.6	0.9
$\text{C}_1-\text{Fe}_1-\text{C}_{11}$	79.6	0.5	$\text{C}_5-\text{C}_6-\text{C}_7$	109.0	0.9
$\text{C}_1-\text{Fe}_1-\text{C}_{12}$	171.8	0.6	$\text{C}_6-\text{C}_7-\text{Fe}_1$	113.5	0.8
$\text{C}_4-\text{Fe}_1-\text{C}_{10}$	173.3	0.5	$\text{Fe}_2-\text{C}_1-\text{C}_2$	75.0	0.7
$\text{C}_4-\text{Fe}_1-\text{C}_{11}$	88.0	0.5	$\text{Fe}_2-\text{C}_2-\text{C}_1$	64.8	0.6
$\text{C}_4-\text{Fe}_1-\text{C}_{12}$	88.5	0.5	$\text{Fe}_2-\text{C}_2-\text{C}_3$	74.3	0.6
$\text{C}_7-\text{Fe}_1-\text{C}_{10}$	90.0	0.5	$\text{Fe}_2-\text{C}_3-\text{C}_2$	67.1	0.6
$\text{C}_7-\text{Fe}_1-\text{C}_{11}$	171.4	0.6	$\text{Fe}_2-\text{C}_6-\text{C}_7$	60.5	0.6
$\text{C}_7-\text{Fe}_1-\text{C}_{12}$	86.0	0.6	$\text{Fe}_2-\text{C}_7-\text{C}_6$	80.6	0.7
$\text{C}_{10}-\text{Fe}_1-\text{C}_{11}$	98.6	0.6	$\text{Fe}_1-\text{C}_1-\text{C}_{13}$	123.7	0.8
$\text{C}_{10}-\text{Fe}_1-\text{C}_{12}$	91.9	0.6	$\text{C}_2-\text{C}_1-\text{C}_{13}$	119.5	1.0
$\text{C}_{11}-\text{Fe}_1-\text{C}_{12}$	94.8	0.6	$\text{C}_1-\text{C}_{13}-\text{C}_{14}$	123.6	1.1
$\text{C}_8-\text{Fe}_2-\text{C}_9$	89.3	0.6	$\text{C}_1-\text{C}_{13}-\text{C}_{18}$	117.8	1.1
$\text{C}_1-\text{Fe}_2-\text{C}_8$	96.0	0.5	$\text{C}_{13}-\text{C}_{14}-\text{C}_{15}$	119.6	1.3
$\text{C}_2-\text{Fe}_2-\text{C}_8$	124.9	0.5	$\text{C}_{14}-\text{C}_{15}-\text{C}_{16}$	119.9	1.5
$\text{C}_3-\text{Fe}_2-\text{C}_8$	163.4	0.5	$\text{C}_{15}-\text{C}_{16}-\text{C}_{17}$	123.2	1.6
$\text{C}_6-\text{Fe}_2-\text{C}_8$	113.9	0.5	$\text{C}_{16}-\text{C}_{17}-\text{C}_{18}$	117.9	1.5
$\text{C}_7-\text{Fe}_2-\text{C}_8$	88.4	0.6	$\text{C}_{17}-\text{C}_{18}-\text{C}_{13}$	120.9	1.2
$\text{C}_1-\text{Fe}_2-\text{C}_9$	135.0	0.6	$\text{C}_{18}-\text{C}_9-\text{C}_{14}$	118.4	1.1
$\text{C}_2-\text{Fe}_2-\text{C}_9$	102.7	0.5	$\text{C}_2-\text{C}_3-\text{C}_{19}$	118.2	0.9
$\text{C}_3-\text{Fe}_2-\text{C}_9$	94.3	0.5	$\text{C}_4-\text{C}_3-\text{C}_{19}$	121.3	0.9
$\text{C}_6-\text{Fe}_2-\text{C}_9$	90.7	0.5	$\text{C}_3-\text{C}_{19}-\text{C}_{20}$	121.4	1.0
$\text{C}_7-\text{Fe}_2-\text{C}_9$	120.8	0.6	$\text{C}_3-\text{C}_{19}-\text{C}_{24}$	119.2	1.0
$\text{C}_1-\text{Fe}_2-\text{C}_7$	104.0	0.5	$\text{C}_{19}-\text{C}_{20}-\text{C}_{21}$	117.6	1.2
$\text{Fe}_1-\text{C}_{10}-\text{O}_4$	176.0	1.2	$\text{C}_{20}-\text{C}_{21}-\text{C}_{22}$	122.0	1.3
$\text{Fe}_1-\text{C}_{11}-\text{O}_5$	176.8	1.2	$\text{C}_{21}-\text{C}_{22}-\text{C}_{23}$	120.2	1.3
$\text{Fe}_1-\text{C}_{12}-\text{O}_6$	176.4	1.3	$\text{C}_{22}-\text{C}_{23}-\text{C}_{24}$	121.0	1.3
$\text{Fe}_2-\text{C}_8-\text{O}_2$	174.4	1.2	$\text{C}_{23}-\text{C}_{24}-\text{C}_{19}$	119.7	1.1
$\text{Fe}_2-\text{C}_9-\text{O}_3$	180.0	1.2	$\text{C}_{24}-\text{C}_{19}-\text{C}_{20}$	119.4	1.0
$\text{Fe}_1-\text{C}_1-\text{Fe}_2$	76.0	0.4	$\text{C}_5-\text{C}_6-\text{C}_{25}$	121.9	0.9
$\text{Fe}_1-\text{C}_7-\text{Fe}_2$	78.0	0.4	$\text{C}_7-\text{C}_6-\text{C}_{25}$	124.9	1.0
$\text{Fe}_1-\text{Fe}_2-\text{C}_1$	54.1	0.3	$\text{C}_6-\text{C}_{25}-\text{C}_{26}$	121.0	1.0
$\text{Fe}_1-\text{Fe}_2-\text{C}_7$	51.7	0.3	$\text{C}_6-\text{C}_{25}-\text{C}_{30}$	121.5	1.0
$\text{Fe}_2-\text{Fe}_1-\text{C}_1$	50.0	0.3	$\text{C}_{25}-\text{C}_{26}-\text{C}_{27}$	120.4	1.2
$\text{Fe}_2-\text{Fe}_1-\text{C}_7$	50.3	0.3	$\text{C}_{26}-\text{C}_{27}-\text{C}_{28}$	123.1	1.4
$\text{Fe}_1-\text{C}_1-\text{C}_2$	110.1	0.8	$\text{C}_{27}-\text{C}_{28}-\text{C}_{29}$	118.2	1.4
$\text{C}_1-\text{C}_2-\text{C}_3$	116.4	1.0	$\text{C}_{28}-\text{C}_{29}-\text{C}_{30}$	118.8	1.3
$\text{C}_2-\text{C}_3-\text{C}_4$	118.5	0.9	$\text{C}_{29}-\text{C}_{30}-\text{C}_{25}$	121.9	1.2
$\text{C}_3-\text{C}_4-\text{Fe}_1$	104.3	0.7	$\text{C}_{30}-\text{C}_{25}-\text{C}_{26}$	117.5	1.1
$\text{Fe}_1-\text{C}_4-\text{C}_5$	102.2	0.7			

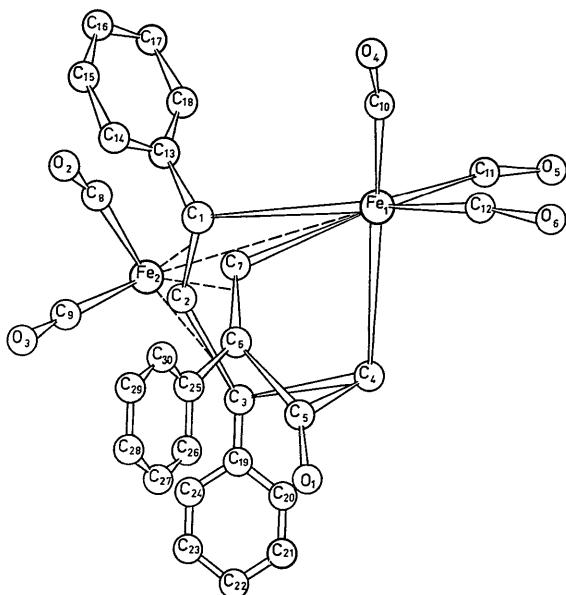
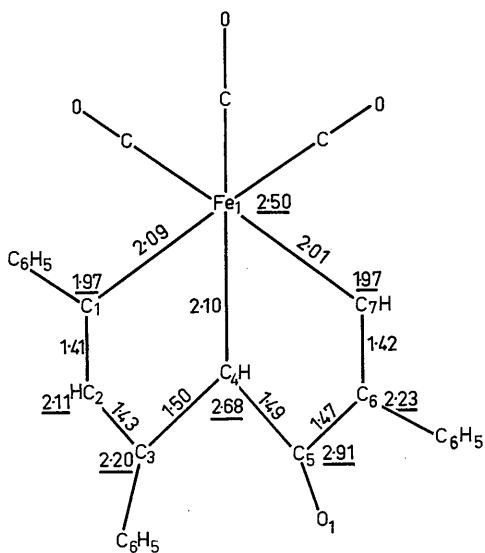
organic system. The second consists of Fe_2 with two carbonyl groups perpendicular to one another. The organic system forms with Fe_1 two five-membered rings having the bond Fe_1-C_4 in common. It is shown in detail in Fig. 2 in which the organic bond lengths are also indicated.

The lengths of the $\text{Fe}-\text{C}$ and $\text{C}-\text{O}$ bonds in the iron carbonyl groups are all normal. The bonds Fe_2-C_1 and Fe_2-C_7 are shorter (1.97 \AA as against 2.12 \AA) than the corresponding bonds in the but-2-yne complex of iron carbonyl hydride, $\text{Fe}_2\text{C}_{12}\text{O}_8\text{H}_8$ (Hock & Mills, 1961). A possible explanation is that the bonds Fe_2-C_1 , Fe_2-C_7 , Fe_1-C_1 and Fe_1-C_7 must be either distorted or shortened in order that the iron–iron distance shall be normal. The angle $\text{C}_1-\text{Fe}_1-\text{C}_7$ is substantially greater than the octahedral angle, the π -bonds Fe_2-C_1 and Fe_2-C_7 are shorter than those in $\text{Fe}_2\text{C}_{12}\text{O}_8\text{H}_8$ and Fe_1-C_7 is shorter than the other $\text{Fe}-\text{C}$ σ -bonds. A reduction in the length of Fe_1-C_1 may be prevented by steric interaction between the phenyl group

attached to C_1 and the carbonyl group $\text{C}_{10}-\text{O}_4$. No explanation can be offered for the fact that the iron–carbon σ -bonds are significantly ($4\frac{1}{2}$ standard deviations for Fe_1-C_7 and 9 standard deviations for Fe_1-C_1 and Fe_1-C_4) longer than in $\text{Fe}_2\text{C}_{12}\text{O}_8\text{H}_8$.

C_3-C_4 and C_4-C_5 do not differ substantially from single bonds while C_5-C_6 is slightly shorter. The three bonds C_1-C_2 , C_2-C_3 and C_6-C_7 are all much shorter and of the same order as those found in $\text{Fe}_2\text{C}_{12}\text{O}_8\text{H}_8$ and in ferrocene (Dunitz, Orgel & Rich, 1956) where there are π -bonds involving iron. The shortest intermolecular contacts are between phenyl and carbonyl groups and are all more than 3.5 \AA .

These results suggest that the two parts of the molecule are linked by a π -bond between C_6-C_7 and Fe_2 , by a three-electron donor system from $\text{C}_1-\text{C}_2-\text{C}_3$ to Fe_2 and by an iron–iron bond. The system $\text{C}_1-\text{C}_2-\text{C}_3$ resembles the allyl group in allylpalladium chloride (Dehm & Chien, 1960) and the group $\text{CHMe}-\text{CH}-\text{CH}_2$ in butadiene cobalt hydrocarbonyl (Moore, Jonassen,

Fig. 1. $\text{Fe}_2(\text{CO})_6(\text{C}_6\text{H}_5\text{C}_2\text{H})_3$.Fig. 2. Bond lengths in $\text{Fe}_2(\text{CO})_6(\text{C}_6\text{H}_5\text{C}_2\text{H})_3$. Values underlined are distances from Fe_2 .

Joyner & Bertrand, 1960) in which nuclear magnetic resonance measurements have shown the allyl carbon-carbon bonds to be equivalent.

The same system exists in cyclopentenylcyclopentadienylnickel(II) (Fischer & Werner, 1961; Jones, Parshall, Pratt & Wilkinson, 1961; Shaw & Sheppard, 1961) where the cyclopentenyl ring corre-

sponds to the ring $\text{C}_1-\text{C}_2-\text{C}_3-\text{C}_4-\text{Fe}_1$ in the present structure.

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