

Organometallic Complexes. VII.* The Structure of the Iron Carbonyl Phenylacetylene Complex, $\text{Fe}_2(\text{CO})_6(\text{C}_6\text{H}_5\text{C}_2\text{H})_3$

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$\text{Fe}_2(\text{CO})_6(\text{C}_6\text{H}_5\text{C}_2\text{H})_3$ crystallizes with four molecules per unit cell in space group $P2_1/n$ with

$$a = 11.963 \pm 0.003, \quad b = 20.442 \pm 0.005, \quad c = 10.326 \pm 0.003 \text{ \AA}; \quad \beta = 93^\circ 24' \pm 5'.$$

The structure has been determined from visually estimated three-dimensional intensity data obtained using Co $K\alpha$ radiation. The iron positions were found from a sharpened Patterson synthesis. Twenty-six of the light atoms were located by use of a minimum function and the remainder from two successive Fourier syntheses. Least-squares refinement led to a final R value of 15%.

The three phenylacetylene residues and one carbonyl group form a seven-membered carbon chain with phenyl groups at the 1, 3 and 6 positions and the carbonyl group in the 5 position. This chain is σ -bonded from the 1, 4 and 7 positions to one iron atom to form two fused five-membered rings with an iron and a carbon atom in common. This iron atom is also bonded to three carbonyl groups and is octahedrally coordinated. The second iron atom is linked to the remaining two carbonyl groups, to the three-electron donor system $\text{C}_1\text{-C}_2\text{-C}_3$ and to $\text{C}_6\text{-C}_7$.

Bond lengths are: Fe-Fe, 2.501 ± 0.003 (e.s.d.); Fe-C (carbonyl) 1.775 ± 0.012 ; Fe-C(σ) 2.006, 2.097, 2.088; Fe-C (three-centre) 1.96-2.23 \AA. All other bond lengths are normal.

Introduction

The reaction between iron dodecacarbonyl and phenylacetylene (Hübel & Braye, 1959) gives a number of organometallic compounds of which $\text{Fe}_2(\text{CO})_6(\text{C}_6\text{H}_5\text{C}_2\text{H})_3$ is of particular interest. Although very stable at ordinary temperatures, it decomposes at its melting point to give 1,3,5-triphenylbenzene. In boiling benzene it is converted to two isomers of triphenyltropone iron tricarbonyl, $\text{Fe}(\text{CO})_4(\text{C}_6\text{H}_5\text{C}_2\text{H})_3$. Recent work (Braye & Hübel, unpublished) has shown that it is also possible to obtain triphenyltropone from $\text{Fe}_2(\text{CO})_6(\text{C}_6\text{H}_5\text{C}_2\text{H})_3$ by treatment with triphenylphosphine. In view of these interesting reactions it seemed worth while to determine the structure of this complex.

Experimental

Crystals of the complex were obtained from a petroleum-ether/benzene solution as described by Hübel & Braye. They are thin black monoclinic plates lying on (010) and bounded by general forms.

Cell dimensions were determined by Farquhar & Lipson's (1946) method from oscillation photographs taken with Co $K\alpha$ radiation

$$(\lambda\alpha_1 = 1.78890, \quad \lambda\alpha_2 = 1.79279 \text{ \AA})$$

and are

$$a = 11.963 \pm 0.003, \quad b = 20.442 \pm 0.005, \\ c = 10.326 \pm 0.003 \text{ \AA}, \quad \beta = 93^\circ 24' \pm 5'.$$

Systematic absences ($h0l$ for $h+l$ odd and $0k0$ for

k odd) were determined from Weissenberg photographs about the b and c axes. The space group is thus $P2_1/n$. This orientation was chosen to make the a and c axes as nearly orthogonal as possible. The observed density is 1.54 ± 0.02 g.cm.⁻³ and that calculated for $Z=4$ is 1.544 g.cm.⁻³.

Multiple-film equi-inclination Weissenberg photographs were taken about the a ($h=0 \rightarrow 4$), b ($k=0$ and 1) and c ($l=0 \rightarrow 6$) axes using crystals $0.18 \times 0.13 \times 0.10$ mm., $0.11 \times 0.08 \times 0.08$ mm., and $0.12 \times 0.09 \times 0.07$ mm., respectively, the shortest dimension being in the b direction. In this way 3415 of the 3685 reflections accessible to Co $K\alpha$ radiation were recorded. Of these 2631 were strong enough to be observed. Intensities were estimated by visual comparison with a set of standard spots prepared from the same crystal. No attempt was made to correct for absorption or for variation in spot shape due to the irregular shapes of the crystals. Allowance was made for the extension and compression of spots on the upper-level Weissenberg photographs following Phillips (1954). Corrections for Lorentz and polarization factors were made by direct computation. Wilson's (1942) method was used to place the intensities on an absolute scale. This method also gave a mean temperature factor of 3.0 \AA^2 .

Determination of the structure

While three-dimensional intensity data were being collected, attempts were made to interpret the c - and a -axis projections. Sharpened Patterson functions were calculated for both these projections. In the

* Part VI. Krüerke, U., Hoogzand, C., & Hübel, W. (1961). *Chem. Ber.* **94**, 2817.

c-axis Patterson projection several more or less consistent sets of possible iron-iron vectors were found. All but one were eliminated by the use of a Buerger (1951) minimum function. A *c*-axis electron-density projection was calculated using signs from the iron-atom contributions alone but could not be interpreted in terms of any conceivable structure. It was impossible unambiguously to determine the *z* coordinates of the iron atoms from the *a*-axis Patterson projection, probably because these coordinates do not differ greatly and are both close to 0.75.

The three-dimensional intensity data were sharpened and then modified by the factor

$$(\sin \theta/\lambda)^4 \exp \{-25.64 (\sin \theta/\lambda)^2\}$$

(Shoemaker *et al.*, 1953) which has a maximum at $\sin \theta = 0.5$ and a very small value at $\sin \theta = 1$. From these modified data a three-dimensional Patterson function was calculated using the *M*-card method* (V. Schomaker, unpublished). The positions of the iron atoms were easily found from the Patterson function. The vectors between one iron atom and the three atoms related to it by the cell symmetry were used as the basic translations in the computation of a minimum function of rank 4. From the minimum function twenty-six light atoms were located in chemically reasonable positions. A retrospective examination showed clearly three more atoms in positions that were *a priori* unreasonable.

These twenty-eight atoms were used to calculate structure factors on an IBM 650 computer. The scattering factors used were taken from Berghuis *et al.* (1955) for C and O, from Freeman & Wood (1959) for Fe and, in the last stages of the work, from McWeeny (1951) for H. The values for Fe were reduced by a dispersion correction of 3.89 electron units calculated from the table given by James (1950). The discrepancy factor *R* was 0.41. A three-dimensional Fourier synthesis showed eight more atoms and a further structure-factor calculation (*R* = 0.38) and Fourier synthesis indicated the positions of the two remaining atoms. A third calculation of structure factors (*R* = 0.28) and Fourier synthesis confirmed these positions. The structure was refined by six applications of the method of least squares using the programme NYXR2 which is a development of NYXR1 (Friedlander, Love & Sayre, 1955) for the IBM 704 computer. The F_o were weighted simply according to the multiplicity with the exception that the non-observed data for which F_o was taken as $(\frac{1}{2}F_{\text{min}}^2)^{\frac{1}{2}}$ were given 20% of this weight. Refinement of an isotropic temperature factor for each atom was introduced in the third cycle. The progress of the refinement is shown in Table 1 where

$$R = \frac{\sum (K|F_o| - |F_c|)}{\sum K|F_o|} \quad \text{and} \quad R' = \frac{\sum w(K|F_o| - |F_c|)^2}{\sum w(K|F_o|)^2}$$

* The master cards were kindly supplied by Dr V. Schomaker.

Table 1. *Progress of the least-squares refinement*

	<i>R</i>	<i>R'</i> 1st weighting	<i>R'</i> 2nd weighting
Parameters from 3rd Fourier synthesis	0.29	—	—
1st least-squares cycle	0.232	120516	—
2nd cycle	0.193	86053	—
3rd cycle, temperature factor refinement intro- duced	0.189	81815	—
4th cycle	0.178	72472	—
5th cycle	0.172	67961	—
6th cycle	0.169	66271	—
6th cycle, non-observed reflections omitted	0.151	64669	313650
7th cycle, hydrogen atoms included	0.145	—	299420
8th cycle	0.149	—	267106
9th cycle	0.152	—	261728

A three-dimensional difference synthesis was calculated using the structure factors calculated during the sixth refinement cycle for reflections of $\sin \theta/\lambda \leq 0.3 \text{ \AA}^{-1}$. There were more or less well-defined positive regions of height ca. 0.4 e.\AA^{-3} in the difference synthesis corresponding to the fifteen phenyl hydrogen atoms. In view of the broadness of some of these peaks these hydrogen atoms were placed on the diagonals of the benzene rings so that $\text{C-H} = 1.05 \text{ \AA}$.

The three remaining hydrogen atoms showed up as relatively sharp peaks from which their positions could be determined. There were about five other peaks of similar size which could not be accounted for by hydrogen atoms and in some cases might be due to anisotropic thermal vibration of phenyl carbon atoms. There were also negative regions (-0.7 e.\AA^{-3}) centred on the iron atoms; these are discussed below.

The refinement was completed by three more least-squares cycles in which hydrogen contributions were included and the non-observed reflections were omitted. The weighting system was changed to $w = 500 p/F^2$ (where *p* is the multiplicity) with a minimum value of 1 and a maximum of 25. This change caused *R* to increase slightly (Table 1) owing to the change in scale factor. The effect of the hydrogen contributions to the structure factors is negligible and the improvement in *R'* during the last three refinement cycles was due mainly to the change in the weighting system. Examination of the values of observed and calculated structure factors from the ninth cycle showed that in most cases where $|F_o| - |F_c|$ was large this quantity was negative, suggesting a systematic error in scale or temperature factors. This was borne out by the negative regions round the iron atoms in the difference Fourier synthesis. It has recently been pointed out (Geller & Durand, 1960) that the programme NYXR2 has a theoretical fault in that it applies the scale factor correction to the observed data which should be kept constant in any least-squares treatment. They too found that this programme gave low values of the temperature factor.

Table 2. Comparison of observed and calculated structure factors
N signifies a negative value

K	F _o	F _c	K	F _o	F _c	K	F _o	F _c	K	F _o	F _c	K	F _o	F _c	K	F _o	F _c	K	F _o	F _c			
00	00	00	06	21	21	05	35	33N	07	17	13N	21	16	18N	04	40	50	11	29	32N	03	K	07
02	101	112N	07	27	22	06	40	39	08	44	48	22	10	12N	05	29	33N	12	46	51	00	39	40
04	44	43N	08	33	32	07	36	34N	09	24	27	06	11	7	06	11	7	13	29	31N	00	39	36
06	75	82N	09	33	32	08	50	52	10	42	39N	07	46	51	07	46	51	14	29	31N	02	23	23N
08	26	25N	10	23	22N	09	107	111	11	30	36N	02	K	01	08	12	7	15	12	11N	03	23	23N
10	76	83N	11	39	42N	10	46	44N	12	30	36N	09	7	6N	09	7	6N	16	30	25	02	26	28
12	38	39	12	28	23N	11	5	2	14	29	25	10	25	21N	10	25	21N	17	28	25	04	11	10N
14	61	72N	13	21	9	12	18	17N	15	5	6N	11	18	13N	18	6	5	18	6	5	05	27	16N
16	14	18N	14	17	13	13	36	34N	17	25	23	13	15	11	19	19	14N	08	20	17N			
18	7	7	15	17	13	14	40	44N	19	7	7	14	16	12	20	15	5	09	20	15N			
20	15	8N	16	20	19	15	22	23	20	11	12	16	12	7N	17	11	7	10	17	12N			
22	8	3N	18	20	19	16	28	31	01	K	06	18	10	8N	03	K	02	12	26	22			
			19	20	19	17	24	23N	02	22	22	19	11	7	04	24	26	13	16	15			
			20	13	14N	18	18	17N	03	51	53N	20	11	12	05	23	21N	08	20	17N			
						19	18	15N	04	51	53N	01	16	17N	06	19	14	09	20	15N			
						20	22	17N	05	21	16	02	10	21	07	19	15	10	17	12N			
						01	K	01	06	21	16	03	12	7	08	17	11	11	16	10			
						02	K	01	07	14	14	04	11	12	09	16	12	12	26	22			
						03	K	01	08	7	4	05	11	12	10	15	11	13	16	15			
						04	K	01	09	14	14	06	15	18	11	15	11	14	16	15			
						05	K	01	10	26	27	07	16	18	12	15	11	15	16	15			
						06	K	01	11	18	17N	08	16	18	13	15	11	16	16	15			
						07	K	01	12	16	16	09	16	18	14	14	11	17	17	15			
						08	K	01	13	16	16	10	17	10N	15	14	11	18	17	15			
						09	K	01	14	16	16	11	15	12N	16	14	11	19	17	15			
						10	K	01	15	15	13N	12	12	12	17	14	11	20	18	15			
						11	K	01	16	16	16	13	12	12	18	14	11						
						12	K	01	17	16	16	14	12	12	19	14	11						
						13	K	01	18	16	16	15	12	12	20	14	11						
						14	K	01	19	16	16	16	12	12									
						15	K	01	20	16	16	17	12	12									
						16	K	01	01	16	17N	18	12	12									
						17	K	01	02	14	14	19	12	12									
						18	K	01	03	14	14	20	12	12									
						19	K	01	04	14	14	21	12	12									
						20	K	01	05	14	14	22	12	12									
									06	14	14	23	12	12									
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												80	12	12									
												81	12	12									
												82	12	12									

Table 2 (cont.)

K	$ F_o $	F_c	K	$ F_o $	F_c	K	$ F_o $	F_c	K	$ F_o $	F_c	K	$ F_o $	F_c	K	$ F_o $	F_c	K	$ F_o $	F_c														
02	30	31	06N K 05	01	23	22	07N K 03	00	27	22N	08N K 01	01	19	12	14	16	11N	03	47	40	05	21	14	03	16	13								
03	65	66N	02	23	24N	01	13	12	02	22	23N	02	28	19	04	39	34	06	33	29	05	19	17N											
04	12	10	03	24	20N	02	16	12N	03	25	12N	03	16	10	05	7	10	06	28	25N	06	17	16	06	11	9N								
05	16	6N	04	38	35	03	25	19N	04	38	42	04	28	25N	07	9	7N	10N K 04	07	9	7N	07	8	10N										
06	7	29	05	37	40N	04	42	44	05	20	18N	05	39	32	09	22	19N	09	22	19N	08	18	18N											
07	26	29	06	32	32N	05	12	12N	06	28	31N	06	28	31N	10	12	7	10N K 05	10	12	7	09	17	15N										
08	18	18N	07	15	15N	06	15	37	07	23	25	07	23	25	11	14	11	11	17	14	11	17	14	10	17	18N								
09	17	15N	08	31	25N	07	46	47	08	35	42N	08	35	42N	12	7	6N	11	17	14	11	17	14	11	17	14	11	17	14					
10	17	18	09	45	44	08	35	34N	09	16	16	09	16	16	13	11	8	10N K 05	10	11	8N	10	11	8N	11	17	14	11	17	14				
11	47	59	10	19	15	11	27	20N	10	23	28	10	23	28	13	11	8	10N K 05	10	11	8N	10	11	8N	11	17	14	11	17	14				
12	17	13	11	19	12N	12	20	13N	11	33	46N	11	33	46N	14	17	20N	11	17	14	10N K 05	10	11	8N	10	11	8N	11	17	14	11	17	14	
13	45	47N	12	17	12N	13	31	24	12	15	12	12	15	12	14	17	20N	11	17	14	10N K 05	10	11	8N	10	11	8N	11	17	14	11	17	14	
14	17	13	13	26	19N	14	36	32	13	17	13	13	17	13	14	17	20N	11	17	14	10N K 05	10	11	8N	10	11	8N	11	17	14	11	17	14	
15	10	15N	14	26	22N	15	19	11N	14	17	12	14	17	12	14	17	20N	11	17	14	10N K 05	10	11	8N	10	11	8N	11	17	14	11	17	14	
16	17	25	15	28	22N	16	19	11N	15	17	12	15	17	12	14	17	20N	11	17	14	10N K 05	10	11	8N	10	11	8N	11	17	14	11	17	14	
17	25	25	16	17	14	17	12	13N	16	17	12	16	17	12	14	17	20N	11	17	14	10N K 05	10	11	8N	10	11	8N	11	17	14	11	17	14	
18	12	8	17	9	6	17	12	13N	17	9	6	17	9	6	14	17	20N	11	17	14	10N K 05	10	11	8N	10	11	8N	11	17	14	11	17	14	
06N K 02	00	20	22N	06N K 06	00	12	6N	07N K 04	01	22	23N	08N K 02	00	28	27N	08N K 08	00	7	12N	09N K 05	00	23	15N	10N K 05	01	18	14	11N K 05	01	18	14			
01	52	62N	01	20	6N	02	16	13	02	17	13	01	18	10	01	7	11N	01	13	12N	02	27	16	02	46	38	02	27	28					
02	11	6	02	8	10	03	20	17	03	10	45N	02	18	10	01	17	11N	02	27	16	03	31	20	03	16	11N	03	16	11N					
03	60	69N	03	19	14N	04	30	27	04	16	13N	03	10	45N	02	34	24N	02	27	16	04	22	12N	04	22	12N	04	22	12N					
04	24	15N	04	29	26N	05	20	17	05	13	5	04	16	13N	01	7	11N	05	34	24N	06	28	25N	05	34	24N	05	12	9N					
05	44	47	05	35	36N	06	23	18N	06	41	44N	05	13	5	08	24	19	06	17	13N	06	28	25N	06	28	25N	06	28	25N					
07	12	2	06	35	36N	07	26	25	07	26	21	06	41	44N	11	17	12	07	25	17N	08	8	6N	07	25	17N	08	8	6N					
08	37	44	07	37	37	08	28	19N	08	39	28	07	26	21	11	17	12	08	24	19	08	8	6N	08	8	6N	08	8	6N					
09	31	33	08	24	20	09	36	34N	09	30	23N	08	39	28	09	17	12	11	17	12	11	17	12	09	10	6	09	10	6					
10	17	13N	09	10	6N	10	32	28	11	40	41N	09	30	23N	09	17	12	12	17	13N	10N K 06	10	19	14N	10N K 06	10	19	14N						
11	22	18N	10	11	14	11	12	27	12	27	27	10	40	41N	09	17	12	12	17	13N	11	17	12	10N K 06	10	19	14N	10N K 06	10	19	14N			
12	24	30N	11	11	11	12	32	17N	13	16	8	11	40	41N	09	17	12	12	17	13N	11	17	12	10N K 06	10	19	14N	10N K 06	10	19	14N			
13	34	38N	12	22	11N	13	23	17N	14	8	4N	12	27	27	09	17	12	12	17	13N	11	17	12	10N K 06	10	19	14N	10N K 06	10	19	14N			
14	23	11	13	9	1	15	19	17N	13	16	8	13	16	8	09	17	12	12	17	13N	11	17	12	10N K 06	10	19	14N	10N K 06	10	19	14N			
15	22	15	14	7	8	17	9	8	14	8	4N	13	16	8	09	17	12	12	17	13N	11	17	12	10N K 06	10	19	14N	10N K 06	10	19	14N			
16	22	15	15	7	8	17	9	8	14	8	4N	13	16	8	09	17	12	12	17	13N	11	17	12	10N K 06	10	19	14N	10N K 06	10	19	14N			
17	19	22N	16	10	14N	16	10	14N	15	19	17N	14	8	4N	13	16	8	09	17	12	12	17	13N	11	17	12	10N K 06	10	19	14N	10N K 06	10	19	14N
06N K 03	01	18	13N	06N K 10	00	22	18N	07N K 05	01	25	22N	08N K 03	01	25	22N	09N K 01	01	13	8	10N K 07	00	10	5N	11N K 01	00	10	5N	12N K 03	02	19	13N			
02	17	12N	00	17	25	02	22	14	03	24	15N	03	24	15N	11	12	9N	11	12	9N	00	19	31N	05	22	20N	05	16	13N	05	16	13N		
03	35	36N	07N K 01	00	56	62	04	26	24N	04	26	24N	04	26	24N	09N K 02	03	27	26	10N K 01	01	22	23N	06	21	17N	06	21	17N	06	21	17N		
04	11	12N	00	17	12N	05	19	13N	05	19	13N	05	19	13N	09	26	22N	10N K 01	03	27	26	07	16	11N	07	16	11N	07	16	11N				
05	45	43	01	24	17N	06	47	35N	06	47	35N	06	47	35N	10	26	22N	09N K 02	04	27	27N	08	8	3	08	8	3	08	8	3				
06	15	7	02	17	12N	08	36	28	08	36	28	08	36	28	11	22	23	10N K 01	04	27	27N	09	5	7N	09	5	7N	09	5	7N				
07	11	7	03	42	41	10	10	7	10	10	7	10	10	7	11	22	23	11N K 02	05	45	46N	10	5	7N	10	5	7N	10	5	7N				
08	42	41	04	32	34N	11	24	13	11	24	13	11	24	13	12	29	24	11N K 02	06	19	14	11	17	18	11	17	18	11	17	18				
09	52	49N	05	37	39	13	24	13N	13	24	13N	13	24	13N	13	24	13N	12	29	24	11N K 02	06	19	14	11	17	18	11	17	18	11	17	18	
10	25	25	06	33	26	16	11	18N	16	11	18N	16	11	18N	16	11	18N	13	24	13N	11N K 02	06	19	14	11	17	18	11	17	18	11	17	18	
11	21	18	07	20	21N	07N K 06	01	15	15	01	15	15	01	15	15	14	16	12	11N K 02	06	19	14	11	17	18	11	17	18	11	17	18			
12	25	20	08	28	28	02	22	14	02	22	14	02	22	14	02	22	14	11N K 02	06	19	14	11	17	18	11	17	18	11	17	18				
13	21	18	09	28	28	03	24	17	03	24	17	03	24	17	03	24	17	11N K 02	06	19	14	11	17	18	11	17	18	11	17	18				
14	27	16	10	28	28	04	26	24N	04	26	24N	04	26	24N	04	26	24N	11N K 02	06	19	14	11	17	18	11	17	18	11	17	18				
15	29	20	11	28	33	05	31	24N	05	31	24N	05	31	24N	05	31	24N	11N K 02	06	19	14	11	17	18	11	17	18	11	17	18				
16	29	20	12	25	24N	06	25	22	06	25	22	06	25	22	06	25																		

Table 5. *Bond angles*

	Angle	Standard deviation		Angle	Standard deviation
$\text{C}_1\text{-Fe}_1\text{-C}_4$	85.1°	0.4	$\text{C}_3\text{-C}_4\text{-C}_5$	107.2°	0.9
$\text{C}_1\text{-Fe}_1\text{-C}_7$	98.6	0.5	$\text{C}_4\text{-C}_5\text{-O}_1$	125.4	1.0
$\text{C}_4\text{-Fe}_1\text{-C}_7$	83.4	0.4	$\text{O}_1\text{-C}_5\text{-C}_6$	121.7	1.0
$\text{C}_1\text{-Fe}_1\text{-C}_{10}$	95.1	0.5	$\text{C}_4\text{-C}_5\text{-C}_6$	112.6	0.9
$\text{C}_1\text{-Fe}_1\text{-C}_{11}$	79.6	0.5	$\text{C}_5\text{-C}_6\text{-C}_7$	109.0	0.9
$\text{C}_1\text{-Fe}_1\text{-C}_{12}$	171.8	0.6	$\text{C}_6\text{-C}_7\text{-Fe}_1$	113.5	0.8
$\text{C}_4\text{-Fe}_1\text{-C}_{10}$	173.3	0.5	$\text{Fe}_2\text{-C}_1\text{-C}_2$	75.0	0.7
$\text{C}_4\text{-Fe}_1\text{-C}_{11}$	88.0	0.5	$\text{Fe}_2\text{-C}_2\text{-C}_1$	64.8	0.6
$\text{C}_4\text{-Fe}_1\text{-C}_{12}$	88.5	0.5	$\text{Fe}_2\text{-C}_2\text{-C}_3$	74.3	0.6
$\text{C}_7\text{-Fe}_1\text{-C}_{10}$	90.0	0.5	$\text{Fe}_2\text{-C}_3\text{-C}_2$	67.1	0.6
$\text{C}_7\text{-Fe}_1\text{-C}_{11}$	171.4	0.6	$\text{Fe}_2\text{-C}_6\text{-C}_7$	60.5	0.6
$\text{C}_7\text{-Fe}_1\text{-C}_{12}$	86.0	0.6	$\text{Fe}_2\text{-C}_7\text{-C}_6$	80.6	0.7
$\text{C}_{10}\text{-Fe}_1\text{-C}_{11}$	98.6	0.6	$\text{Fe}_1\text{-C}_1\text{-C}_{13}$	123.7	0.8
$\text{C}_{10}\text{-Fe}_1\text{-C}_{12}$	91.9	0.6	$\text{C}_5\text{-C}_1\text{-C}_{13}$	119.5	1.0
$\text{C}_{11}\text{-Fe}_1\text{-C}_{12}$	94.8	0.6	$\text{C}_1\text{-C}_{13}\text{-C}_{14}$	123.6	1.1
$\text{C}_5\text{-Fe}_2\text{-C}_9$	89.3	0.6	$\text{C}_1\text{-C}_{13}\text{-C}_{18}$	117.8	1.1
$\text{C}_1\text{-Fe}_2\text{-C}_8$	96.0	0.5	$\text{C}_{13}\text{-C}_{14}\text{-C}_{15}$	119.6	1.3
$\text{C}_2\text{-Fe}_2\text{-C}_8$	124.9	0.5	$\text{C}_{14}\text{-C}_{15}\text{-C}_{16}$	119.9	1.5
$\text{C}_3\text{-Fe}_2\text{-C}_8$	163.4	0.5	$\text{C}_{15}\text{-C}_{16}\text{-C}_{17}$	123.2	1.6
$\text{C}_6\text{-Fe}_2\text{-C}_8$	113.9	0.5	$\text{C}_{16}\text{-C}_{17}\text{-C}_{18}$	117.9	1.5
$\text{C}_7\text{-Fe}_2\text{-C}_8$	88.4	0.6	$\text{C}_{17}\text{-C}_{18}\text{-C}_{13}$	120.9	1.2
$\text{C}_1\text{-Fe}_2\text{-C}_9$	135.0	0.6	$\text{C}_{18}\text{-C}_{13}\text{-C}_{14}$	118.4	1.1
$\text{C}_2\text{-Fe}_2\text{-C}_9$	102.7	0.5	$\text{C}_2\text{-C}_9\text{-C}_{19}$	118.2	0.9
$\text{C}_3\text{-Fe}_2\text{-C}_9$	94.3	0.5	$\text{C}_4\text{-C}_3\text{-C}_{19}$	121.3	0.9
$\text{C}_6\text{-Fe}_2\text{-C}_9$	90.7	0.5	$\text{C}_3\text{-C}_{19}\text{-C}_{20}$	121.4	1.0
$\text{C}_7\text{-Fe}_2\text{-C}_9$	120.8	0.6	$\text{C}_3\text{-C}_{19}\text{-C}_{24}$	119.2	1.0
$\text{C}_1\text{-Fe}_2\text{-C}_7$	104.0	0.5	$\text{C}_{19}\text{-C}_{20}\text{-C}_{21}$	117.6	1.2
$\text{Fe}_1\text{-C}_{10}\text{-O}_4$	176.0	1.2	$\text{C}_{20}\text{-C}_{21}\text{-C}_{22}$	122.0	1.3
$\text{Fe}_1\text{-C}_{11}\text{-O}_5$	176.8	1.2	$\text{C}_{21}\text{-C}_{22}\text{-C}_{23}$	120.2	1.3
$\text{Fe}_1\text{-C}_{12}\text{-O}_6$	176.4	1.3	$\text{C}_{22}\text{-C}_{23}\text{-C}_{24}$	121.0	1.3
$\text{Fe}_2\text{-C}_8\text{-O}_2$	174.4	1.2	$\text{C}_{23}\text{-C}_{24}\text{-C}_{19}$	119.7	1.1
$\text{Fe}_2\text{-C}_9\text{-O}_3$	180.0	1.2	$\text{C}_{24}\text{-C}_{19}\text{-C}_{20}$	119.4	1.0
$\text{Fe}_1\text{-C}_1\text{-Fe}_2$	76.0	0.4	$\text{C}_5\text{-C}_6\text{-C}_{25}$	121.9	0.9
$\text{Fe}_1\text{-C}_7\text{-Fe}_2$	78.0	0.4	$\text{C}_7\text{-C}_6\text{-C}_{25}$	124.9	1.0
$\text{Fe}_1\text{-Fe}_2\text{-C}_1$	54.1	0.3	$\text{C}_5\text{-C}_{25}\text{-C}_{26}$	121.0	1.0
$\text{Fe}_1\text{-Fe}_2\text{-C}_7$	51.7	0.3	$\text{C}_6\text{-C}_{25}\text{-C}_{30}$	121.5	1.0
$\text{Fe}_2\text{-Fe}_1\text{-C}_1$	50.0	0.3	$\text{C}_{25}\text{-C}_{26}\text{-C}_{27}$	120.4	1.2
$\text{Fe}_2\text{-Fe}_1\text{-C}_7$	50.3	0.3	$\text{C}_{26}\text{-C}_{27}\text{-C}_{28}$	123.1	1.4
$\text{Fe}_1\text{-C}_1\text{-C}_2$	110.1	0.8	$\text{C}_{27}\text{-C}_{28}\text{-C}_{29}$	118.2	1.4
$\text{C}_1\text{-C}_2\text{-C}_3$	116.4	1.0	$\text{C}_{28}\text{-C}_{29}\text{-C}_{30}$	118.8	1.3
$\text{C}_2\text{-C}_3\text{-C}_4$	118.5	0.9	$\text{C}_{29}\text{-C}_{30}\text{-C}_{25}$	121.9	1.2
$\text{C}_5\text{-C}_4\text{-Fe}_1$	104.3	0.7	$\text{C}_{30}\text{-C}_{25}\text{-C}_{26}$	117.5	1.1
$\text{Fe}_1\text{-C}_4\text{-C}_5$	102.2	0.7			

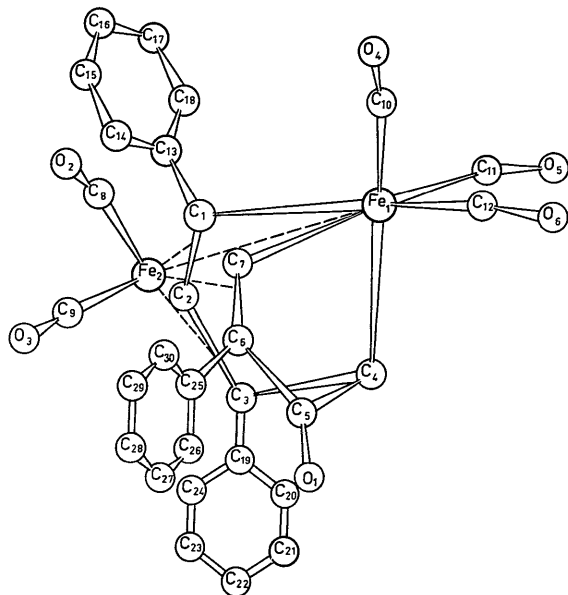
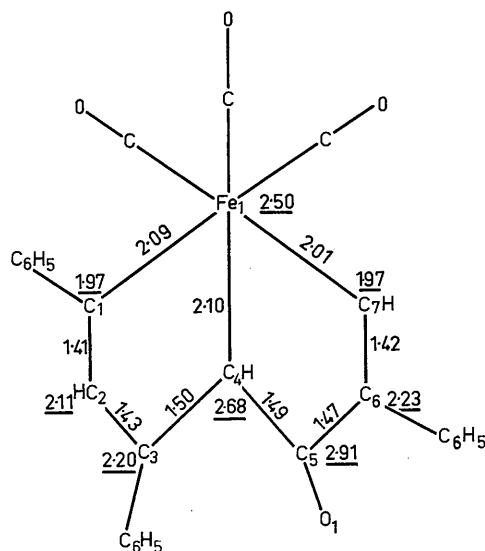
organic system. The second consists of Fe_2 with two carbonyl groups perpendicular to one another. The organic system forms with Fe_1 two five-membered rings having the bond $\text{Fe}_1\text{-C}_4$ in common. It is shown in detail in Fig. 2 in which the organic bond lengths are also indicated.

The lengths of the Fe-C and C-O bonds in the iron carbonyl groups are all normal. The bonds $\text{Fe}_2\text{-C}_1$ and $\text{Fe}_2\text{-C}_7$ are shorter (1.97 Å as against 2.12 Å) than the corresponding bonds in the but-2-yne complex of iron carbonyl hydride, $\text{Fe}_2\text{C}_{12}\text{O}_8\text{H}_8$ (Hock & Mills, 1961). A possible explanation is that the bonds $\text{Fe}_2\text{-C}_1$, $\text{Fe}_2\text{-C}_7$, $\text{Fe}_1\text{-C}_1$ and $\text{Fe}_1\text{-C}_7$ must be either distorted or shortened in order that the iron-iron distance shall be normal. The angle $\text{C}_1\text{-Fe}_1\text{-C}_7$ is substantially greater than the octahedral angle, the π -bonds $\text{Fe}_2\text{-C}_1$ and $\text{Fe}_2\text{-C}_7$ are shorter than those in $\text{Fe}_2\text{C}_{12}\text{O}_8\text{H}_8$ and $\text{Fe}_1\text{-C}_7$ is shorter than the other Fe-C σ -bonds. A reduction in the length of $\text{Fe}_1\text{-C}_1$ may be prevented by steric interaction between the phenyl group

attached to C_1 and the carbonyl group $\text{C}_{10}\text{-O}_4$. No explanation can be offered for the fact that the iron-carbon σ -bonds are significantly ($4\frac{1}{2}$ standard deviations for $\text{Fe}_1\text{-C}_7$ and 9 standard deviations for $\text{Fe}_1\text{-C}_1$ and $\text{Fe}_1\text{-C}_4$) longer than in $\text{Fe}_2\text{C}_{12}\text{O}_8\text{H}_8$.

$\text{C}_3\text{-C}_4$ and $\text{C}_4\text{-C}_5$ do not differ substantially from single bonds while $\text{C}_5\text{-C}_6$ is slightly shorter. The three bonds $\text{C}_1\text{-C}_2$, $\text{C}_2\text{-C}_3$ and $\text{C}_6\text{-C}_7$ are all much shorter and of the same order as those found in $\text{Fe}_2\text{C}_{12}\text{O}_8\text{H}_8$ and in ferrocene (Dunitz, Orgel & Rich, 1956) where there are π -bonds involving iron. The shortest intermolecular contacts are between phenyl and carbonyl groups and are all more than 3.5 Å.

These results suggest that the two parts of the molecule are linked by a π -bond between $\text{C}_6\text{-C}_7$ and Fe_2 , by a three-electron donor system from $\text{C}_1\text{-C}_2\text{-C}_3$ to Fe_2 and by an iron-iron bond. The system $\text{C}_1\text{-C}_2\text{-C}_3$ resembles the allyl group in allylpalladium chloride (Dehm & Chien, 1960) and the group CHMe-CH-CH_2 in butadiene cobalt hydrocarbonyl (Moore, Jonassen,

Fig. 1. $\text{Fe}_2(\text{CO})_6(\text{C}_6\text{H}_5\text{C}_2\text{H})_3$.Fig. 2. Bond lengths in $\text{Fe}_2(\text{CO})_6(\text{C}_6\text{H}_5\text{C}_2\text{H})_3$.
Values underlined are distances from Fe_2 .

Joyner & Bertrand, 1960) in which nuclear magnetic resonance measurements have shown the allyl carbon-carbon bonds to be equivalent.

The same system exists in cyclopentenylcyclopentadienylnickel(II) (Fischer & Werner, 1961; Jones, Parshall, Pratt & Wilkinson, 1961; Shaw & Sheppard, 1961) where the cyclopentenyl ring corre-

sponds to the ring $\text{C}_1\text{-C}_2\text{-C}_3\text{-C}_4\text{-Fe}_1$ in the present structure.

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